



HAL
open science

Physico-Chemical Heterogeneity of Organic-Rich Sediments in the Rifle Aquifer, CO: Impact on Uranium Biogeochemistry

Noemie Janot, Juan S Lezama Pacheco, Don Q Pham, Timothy O'brien, Debra Hausladen, Vincent Noël, Florent Lallier, Kate Maher, Scott Fendorf, Kenneth Williams, et al.

► To cite this version:

Noemie Janot, Juan S Lezama Pacheco, Don Q Pham, Timothy O'brien, Debra Hausladen, et al.. Physico-Chemical Heterogeneity of Organic-Rich Sediments in the Rifle Aquifer, CO: Impact on Uranium Biogeochemistry. *Environmental Science and Technology*, 2016, 50 (1), pp.46-53. <10.1021/acs.est.5b03208>. <hal-02535730>

HAL Id: hal-02535730

<https://hal.inrae.fr/hal-02535730v1>

Submitted on 6 Dec 2023

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire HAL, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.



HAL Authorization

Lawrence Berkeley National Laboratory

Climate & Ecosystems

Title

Physico-Chemical Heterogeneity of Organic-Rich Sediments in the Rifle Aquifer, CO: Impact on Uranium Biogeochemistry

Permalink

<https://escholarship.org/uc/item/78r5k9h9>

Journal

Environmental Science and Technology, 50(1)

ISSN

0013-936X

Authors

Janot, Noémie
Pacheco, Juan S Lezama
Pham, Don Q
[et al.](#)

Publication Date

2016-01-05

DOI

10.1021/acs.est.5b03208

Peer reviewed

Physico-Chemical Heterogeneity of Organic-Rich Sediments in the Rifle Aquifer, CO: Impact on Uranium Biogeochemistry

[Noémie Janot](#)[†], [Juan S. Lezama Pacheco](#)^{†‡}, [Don Q. Pham](#)[†], [Timothy M. O'Brien](#)[†], [Debra Hausladen](#)[‡], [Vincent Noël](#)[†], [Florent Lallier](#)[§], [Kate Maher](#)[¶], [Scott Fendorf](#)[‡], [Kenneth H. Williams](#)[⊥], [Philip E. Long](#)[⊥], and [John R. Bargar](#)^{†‡}

[†] Stanford Synchrotron Radiation Lightsource, SLAC National Accelerator Laboratory, Menlo Park, California 94025, United States

[‡] Department of Environmental Earth System Science, Stanford University, Stanford, California 94305, United States

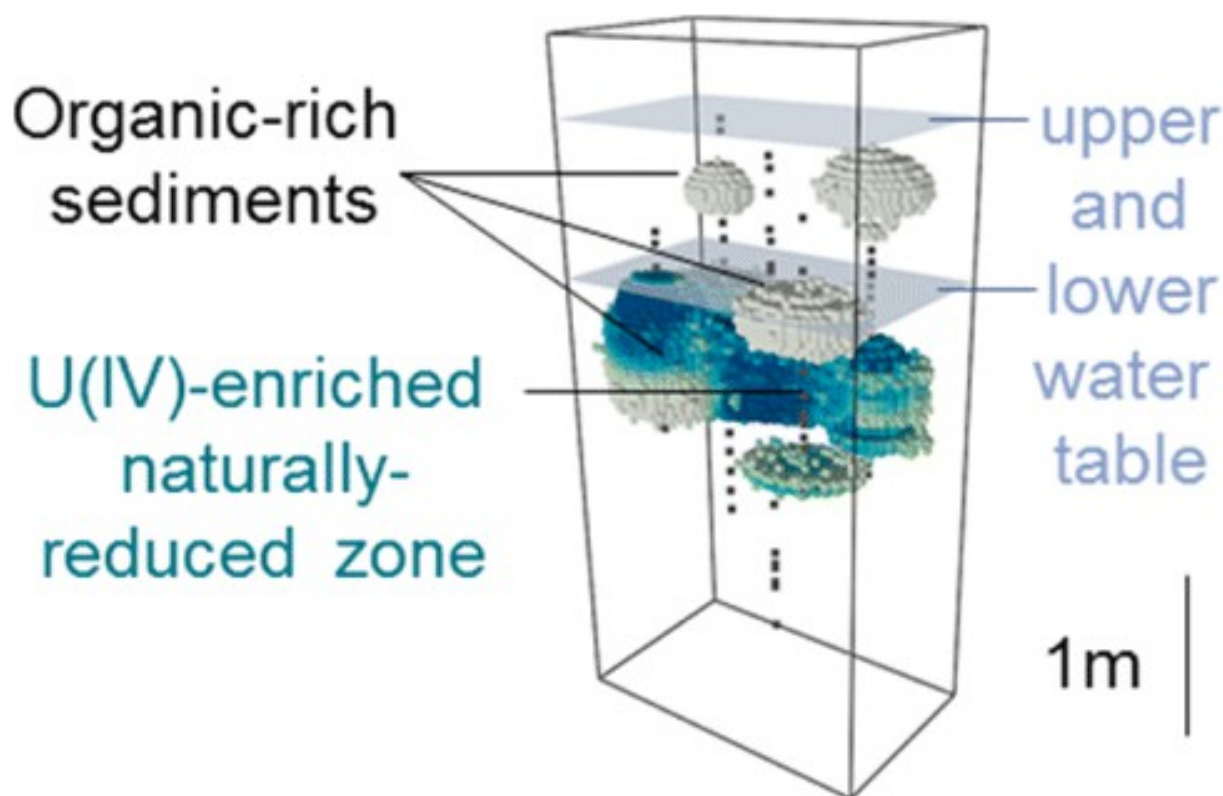
[§] Université de Lorraine, Vandœuvre-lès-Nancy, F-54501, France

[¶] Department of Geological Sciences, Stanford University, Stanford, California 94305, United States

[⊥] Earth Sciences Division, Lawrence Berkeley National Laboratory, Berkeley, California 94720, United States

Environ. Sci. Technol., **2016**, *50* (1), pp 46–53

DOI: 10.1021/acs.est.5b03208



The Rifle alluvial aquifer along the Colorado River in west central Colorado contains fine-grained, diffusion-limited sediment lenses that are substantially enriched in organic carbon and sulfides, as well as uranium, from previous milling operations. These naturally reduced zones (NRZs) coincide spatially with a persistent uranium groundwater plume. There is concern that uranium release from

NRZs is contributing to plume persistence or will do so in the future. To better define the physical extent, heterogeneity and biogeochemistry of these NRZs, we investigated sediment cores from five neighboring wells. The main NRZ body exhibited uranium concentrations up to 100 mg/kg U as U(IV) and contains ca. 286 g of U in total. Uranium accumulated only in areas where organic carbon and reduced sulfur (as iron sulfides) were present, emphasizing the importance of sulfate-reducing conditions to uranium retention and the essential role of organic matter. NRZs further exhibited centimeter-scale variations in both redox status and particle size. Mackinawite, greigite, pyrite and sulfate coexist in the sediments, indicating that dynamic redox cycling occurs within NRZs and that their internal portions can be seasonally oxidized. We show that oxidative U(VI) release to the aquifer has the potential to sustain a groundwater contaminant plume for centuries. NRZs, known to exist in other uranium-contaminated aquifers, may be regionally important to uranium persistence.

Introduction

At U.S. Department of Energy (DOE) legacy uranium ore processing sites across the Colorado River Basin (CRB) (e.g., Rifle and Naturita, CO(1, 2)), contaminated surface materials have been removed but aquifer contamination has subsequently persisted for decades. The mechanisms of plume persistence are unknown, but suspicion has fallen on slow pervasive uranium release from fine-grained, sulfidic sediment lenses (i.e., naturally reduced zones, NRZs) that are enriched in uranium. (2) In order for NRZs to contribute to persistent plumes, they must contain sufficient uranium inventories to sustain releases and exhibit both physical and biogeochemical conditions conducive for oxidation. Reduced U(IV) is relatively immobile, providing a mechanism for its accumulation in reduced sediments, whereas oxidized U(VI) is soluble and highly mobile in Rifle groundwater. (3) Consequently, redox zonation within sediments controls uranium speciation, mobility, and concentration. NRZs have been observed in multiple uranium-contaminated aquifers within the upper CRB and exhibit uranium concentrations (up to 48 mg_U/kg(4, 5)) significantly higher than found in the bulk sediments in the surrounding aquifer (ca. 1 mg_U/kg). A robust understanding of redox heterogeneity, biogeochemical processes, and transport processes in NRZs is necessary in order to assess the importance of NRZs to uranium plume persistence in these floodplain aquifers. Campbell et al. (2012) concluded that the lower permeability of NRZs observed at the former uranium ore processing site at Rifle, CO, creates biogeochemical conditions favorable for microbial reduction of iron, sulfur, and uranium. The spatial resolution of this former study, however, was

relatively coarse (1 m vertical intervals). Detailed observations by our group during coring activities at the Rifle site in 2011 showed that NRZs are layered at the cm scale and exhibit textural and compositional heterogeneity over length scales down to ca. 1 mm, a finer spatial scale than previous reports. Therefore, more detailed qualitative and quantitative information on the physical and biogeochemical attributes of NRZ sediments are needed.

The purposes of this study were (i) to map redox zoning of Rifle sediments across NRZs at scales fine enough to help understand their impact on uranium biogeochemistry and (ii) to better estimate NRZs dimensions as well as the sizes of the pools of reduced uranium, sulfide, and organic carbon. Such information supports the development of improved geochemical models of uranium mobility and floodplain biogeochemical function. To achieve these objectives, five cores were drilled through a previously observed NRZ,[\(4\)](#) sampled at a vertical resolution of about 10 cm, and analyzed using a suite of techniques.

Materials and Methods

Site Description

The Rifle field research site is a DOE-managed former uranium and vanadium ore processing mill facility near the town of Rifle, CO, as described elsewhere.[\(2, 3, 6-8\)](#) Between 1992 and 1996, uranium-rich mill tailings and contaminated soil were removed to a depth of ~1.6 m and replaced with clean compacted clay fill as part of the Uranium Mill Tailings Remedial Action (UMTRA). However, uranium levels in groundwater have not decreased significantly since 1999, and some estimates suggest they will remain above regulatory limits (0.044 mg_U/L) for tens if not hundreds of years.[\(2, 3, 9\)](#)

Sediment Collection

Five sediment cores (150 mm diameter), labeled JB-01 to JB-05, were collected in proximity to the pre-existing well P-103 in a known NRZ[\(4\)](#) (Figure S1 in [Supporting Information](#)). Sediments were collected from the depth of the base of the clean sediment cap (1.8 m below ground surface (bgs)) down to bedrock (ca. 5.8 m bgs) across multiple wells at relatively fine vertical resolution (~10 cm). Sediments from the five JB wells were excavated using a rotasonic drill, encapsulated immediately into N₂-purged low density polyethylene sleeves, each approximately 0.7 m in length, and moved into the on-site research laboratory on July 22–24, 2011. Pebbles and cobbles larger than about 10 mm, which accounted for ca. 90% of the mass, were removed, and the remaining sediments were immediately placed into containers. Sediments that were visibly reduced, as identified by the

presence of dark coloration, fine grain sizes and sulfidic smell, were collected at ca. 10 cm intervals and rapidly packed into glass jars. Voids were compacted to minimize entrapment of air, the shallow headspaces were immediately covered with groundwater to the top of the jar to exclude air, and metal canning jar lids were affixed. Sand-size and gravelly sediments that did not visually display evidence of reduction were sealed into N₂-purged double aluminized Mylar bags with oxygen absorbers. Sediments were shipped within 72 h of collection to the SLAC National Accelerator Laboratory where they were stored at 3 °C until analysis. During all subsequent sample manipulation and preparation steps, glass jars were opened in an anaerobic chamber to avoid oxidation of the material. Subsamples taken after homogenization of the material within the glass jar were then either dried inside the anaerobic chamber before spectroscopic analysis or air-dried for total elemental composition analysis. The samples analyzed in this study represent the <2 mm size fraction of these sediments.

Elemental Composition Analysis

For each sample, ~20 g of sediment were subsampled (after homogenization of the jar material) and air-dried. After drying, samples were sieved and the <2 mm size fraction was ground using a shatterbox with WC vessels for 2 min.

X-ray fluorescence (XRF) analyses were conducted with a Spectro Xepos instrument on ~4 g of powdered sediments at Environmental Measurement I (EM-1) laboratory at Stanford University, CA. Solid phase carbon (total and organic) was measured using an elemental analyzer (EA) Carbo Erba NA1500. Around 30 mg of ground sample was weighed in tin cups. For organic carbon (OC) measurements, samples were acidified with 2 drops of 8% H₃PO₄ and allowed to degas overnight in order to remove inorganic carbonates before analysis.

A subset of selected samples were dried and ~100 g of the <2 mm fraction was sieved into 5 fractions using a standard series of mesh sizes for 10 min. Sieving was performed using a set of 3" stainless steel and brass nested sieves (53 µm, 150 µm, 500 µm, 1 mm, 2 mm) by hand shaking. Size fractions were collected and analyzed for elemental composition.

Carbon Dating

Subsets of sediments were shipped to Beta Analytics for radiocarbon dating measurements on fine root, woody fragments and plant material. Radiocarbon dating was determined by the accelerator mass spectrometry (AMS) method measuring directly the amount of ¹⁴C in the sample.

Spectroscopy Analysis

Sulfur K-edge X-ray absorption spectroscopy (XAS) were measured from wet anoxic sediments to map the distribution of sulfur oxidation states. Uranium XAS was used to determine the oxidation

state and molecular structure around uranium. Data collection and analysis operations are described in the [Supporting Information](#).

Results

Sediment Texture and Color

Textures and colors of cored sediments were highly variable. A summary of sediment descriptions is provided in [Table S1](#). Most of the cores were comprised of oblate spheroidal siliciclastic cobbles larger than 5 cm diameter (with major axis ≤ 20 cm), pebbles, gravel, and sand exhibiting a light brown color. Fine-grained sediment was commonly present as thin layers interleaved between and infilling interstices created by cobbles ([Figure 1](#)). Within fine-grained sediments, redox zonation was apparent from brown-to-black color contrasts defined by sharp boundaries, indicating local areas of sulfate reduction and precipitation of dark-colored metal sulfides. These darker fine-grained materials were observed in lenses at different depths in the different cores: at 4.0–4.5 m bgs in JB-01, at 4.0–4.3 m bgs and 5.5 m bgs in JB-02, 3.8–4.6 m bgs in JB-04, at 3.5–3.7 and 3.8–4.1 m bgs in JB-05, whereas none were observed in JB-03.



Figure 1. Pictures of Rifle sediments (left, a–c) and nearby contemporaneous river bank deposits (right, d–e, Roaring Fork river, Basalt, CO). Rifle sediments from well JB-05 (a) contain abundant large cobbles with reduced clay/silt in between. Sediment details are shown for Rifle organic-rich sediments (JB-05, b–c) and stream bank deposits (e). Note the presence of fine roots and sharp redox boundaries in both NRZs and river-bank deposits (b and e, respectively).

Vertical and Horizontal Distributions of Trace and Minor Elements

Sulfur, uranium and OC concentration profiles in the five JB cores are shown in [Figure 2](#). Well-defined layers of high concentrations were present, with sharp variations at the centimeter scale. For example, in well JB-02, U and OC concentrations increase from 13 mg/kg and 0.3 wt %, respectively, to 55 mg/kg and 1.0 wt % between 4.4 and 4.5 m bgs.



Figure 2. Vertical variations of sulfur (diamonds), uranium (squares) and organic carbon (triangles) concentrations in sediment cores JB-01 to JB-05 (left to right). Samples analyzed for radiocarbon dating and results are circled in red.

Zones of elevated sulfur, OC, and uranium exhibit variable thicknesses and were found at different depths in the different JB wells. In well JB-01, we observed a zone enriched in U/S/OC at 3.5–4.0 m bgs. In well JB-02, this layer occurred at 4.0–4.5 m bgs (which coincides with the black fine-grained sediments observed during sample collection, [Table S1](#)) and contained up to 102 mg_U/kg (4.3 m bgs). This value is higher than any previously reported natural concentrations of uranium reported for the upper CRB to our knowledge: acid-extractable concentrations up to 64 mg_U/kg were reported at the Naturita site([1](#)) whereas at Rifle maximum concentrations of 19 mg_U/kg([10](#)) and 48 mg_U/kg([4](#)) were found in wells D-08 and P-103 (at 4.9 and 3.1 m bgs, respectively). Sediments from JB-01 and JB-02 show up to 1.7% OC, which exceeds previously reported concentrations of 1% (well P-103([4](#))). In well JB-04, sediments enriched in U/S/OC were present at 3.8–4.6 m bgs, and exhibited variable concentrations. This variability was also observed at the macroscopic scale ([Table S1](#)); root material embedded in mottled light and dark sulfidic sediments were present in these samples. In wells JB-03 and JB-05, sediments were less enriched in uranium (≤ 12 mg/kg and ≤ 3 mg/kg in JB-03 and -05, respectively) and in OC ($\leq 0.3\%$ and $\leq 0.9\%$ in JB-03 and -05, respectively). Total iron concentration depth profiles do not show marked variations ([Figure S2](#)), but reduced sediments are slightly enriched in Fe, with concentrations up to 3.3% in JB-02 NRZ when they are at 2.0–2.5% in the background aquifer sediments, in agreement with previous observations. ([4](#))Manganese profiles show similar variations as Fe ([Figure S2](#)), with concentrations between 200 and 600 mg/kg.

Depositional Age

Radiocarbon dating was performed on detrital organic material (roots, twigs) found in a subset of samples from various wells and depths. Organic carbon ages range from 110 to 180 ± 30 yr BP (calibrated to years before 1950, [Figure 2](#)). These dates provide maximum ages for the time of deposition of overlying sediments, assuming that the roots and twigs analyzed have not equilibrated with modern carbon introduced by microbes or modern DOC.

Redox Variability with Depth

[Figure 3](#) shows the sulfur K-edge XANES spectra of samples from well JB-02, from 3.2 to 4.5 m bgs. The loss of sulfate as a percentage of total S species with depth is clearly seen, the characteristic 2482 eV peak diminishing in the spectra from sediments below 4.0 m bgs. In deeper samples, spectra show peaks between 2470 and 2475 eV, characteristic of inorganic sulfides ([Figure S3](#)). No organic forms of sulfur, whose spectra show peaks between 2473 and 2481 eV, ([11](#)) were detected.

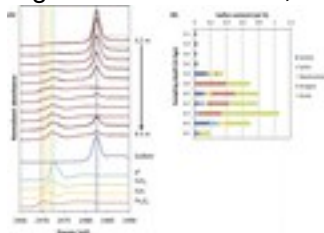


Figure 3. (A) Sulfur XANES spectra of natural sediments as a function of depth in well JB-02 (black lines), together with the reference compounds used for linear combination fitting (LCF, colored lines) and fit results (red lines). (B) Decomposition of sulfur speciation estimated from LCF results.

Linear combination fitting (LCF) results are shown in [Figure 3](#) (details in [Figures S4 and S5](#)) and parameters are summarized in [Table S2](#). The presence of iron sulfides (FeS , FeS_2 , Fe_3S_4) correlates with increasing concentrations of total sulfur species. At the interface with the aquifer (around 3.9 m bgs), reduced sulfur is mainly present in the form of monosulfides (mackinawite and greigite). Greigite was not observed in a previous study using Mössbauer spectroscopy on similar samples, ([4](#)) but its absence from the samples could be explained by the spatial variability of its abundance. In contrast, pyrite is the most abundant sulfide in the interior of the NRZ. Detailed results of the LCF are shown in [Figure S6](#) for two of these samples: one from the aquifer (3.4 m bgs), one from the NRZ (4.3 m bgs).

Sediment Particle Size-Dependence of Sulfur, OC, and Uranium

Sieving of JB-02 samples showed that fine particles ($<53 \mu\text{m}$) are more abundant in NRZs (4.2–4.4 m bgs) than in the surrounding aquifer ([Figure S7](#)). The main difference between NRZs and aquifer sediments resides in the proportion of $<150 \mu\text{m}$ fraction, in agreement with Campbell et al. (2012): ([4](#)) in all cases this fraction represents more than 25% of the $<2000 \mu\text{m}$ fraction, and its proportion reaches 69% in the 4.2 m bgs sample.

Sulfur is most abundant in the smallest size fractions in all analyzed samples ([Table 1](#)), but this is not the case for uranium. In the nonreduced samples (3.4 and 3.6 m bgs), we observed that uranium and OC are most abundant in the smallest size fraction. In NRZs, however, uranium is most abundant in the fraction with the most OC, that is, in the 150–500 μm fraction, in two of the three samples. In these reduced samples, uranium concentration is lowest in the 53–150 μm fraction. The

a priori expectation was that more uranium would be present in the fine- versus the coarse-grained fraction, as reported previously in similar samples.⁽⁴⁾ Differences in our results illustrate the substantial heterogeneity in aquifer composition and sampled material.

Table 1. Uranium, Sulfur and Organic Carbon Contents in Size Fractions of JB-02 Samples

			U (mg/kg)	S (%)	OC (%)
aquifer	3.4 m	bulk	<0.6	0.04	0.1
		< 53 μm	5	0.13	1.0
		53–150 μm	2	0.08	0.4
		150–500 μm	<0.4	0.04	0.2
	3.6 m	bulk	<0.6	0.06	0.2
		< 53 μm			
		53–150 μm	6	0.18	0.8
		150–500 μm	2	0.10	0.4

			U (mg/kg)	S (%)	OC (%)
NRZ	4.2 m	bulk	74	0.79	1.5
		< 53 μm	51	1.17	1.0
		53–150 μm	46	0.92	1.0
		150–500 μm	80	0.84	2.0
	4.3 m	bulk	102	1.06	1.6
		< 53 μm	101	1.13	1.6
		53–150 μm	84	0.90	1.3
		150–500 μm	110	0.98	1.8

		U (mg/kg)	S (%)	OC (%)
4.4 m	bulk	56	0.70	0.9
	< 53 μm	85	1.17	1.2
	53–150 μm	69	0.95	0.9
	150–500 μm	79	0.74	1.3

Uranium Oxidation State and Local Structure

Uranium L_{II} -edge XANES spectra from JB-01 and JB-02 samples are presented in [Figure S8](#). LCF of XANES spectra using uraninite as U(IV) standard and autunite as U(VI) standard gave $100 \pm 10\%$ U(IV) in all the JB-02 samples, and $86 \pm 10\%$ U(IV) in the JB-01 sample. EXAFS and Fourier Transform of uranium in two JB-02 NRZ samples, 4.2 and 4.3 m bgs, are shown in [Figure S9](#). They are similar to those of U(IV) in a P-103 sample. (4) Results of shell-by-shell fitting are summarized in [Table S3](#). Both spectra show a small second shell frequency around 3.85 Å corresponding to U–U pair correlation, implying minor amounts of UO_2 in these sediments, but no U–P pair correlation was detected.

Discussion

Compositional Heterogeneity of the Aquifer

Analysis of the samples shows that the physicochemical and redox conditions of the aquifer differ vertically on a cm scale and laterally on a scale of less than 2 m ([Figure S1](#)). We operationally define NRZs as samples with more than 0.35% S. Within the five cores studied here, naturally reduced sediments represent 20% of total analyzed samples and they contain 90% of the total sediment

uranium inventory and 56% of the total sediment OC inventory: these structures accumulate large inventories of these elements. In addition, homogenization of sampled material may result in dilution of smaller-scale structures: pictures of JB-05 samples taken during excavation reveal the presence of locally black, reduced sediments ([Figure 1](#)) that do not coincide with elevated sulfur concentrations. NRZs identified in this study are thus structures larger than the subdecimeter scale, but it is important to note that smaller structures are also existent.

JB-01, JB-02, and JB-04 each shows a well-defined OC-rich lens with sharp margins ([Figure 2](#)), with a top elevation decreasing from 3.5 m to 4.0 m bgs from JB-01 to JB-04 along a SW-NE axis ([Figure S1](#)). These three cores seem to belong to a single NRZ structure that thickens between JB-01 and JB-04 and likely extends beyond the boundaries of the well cluster. In comparison, JB-03 and JB-05 exhibit much lower concentrations of OC distributed more diffusely over a broader range of elevations than in the other three cores. These two wells lie on the periphery of the JB-01-02-04 axis, suggesting that they represent the margins of an elongate NRZ body ([Figure 5](#)).

Sulfur profiles show well-defined enrichment zones within the main JB-01/02/04 NRZ and are coincident with the OC enrichment profiles ([Figure 2](#)). This suggests that OC content helps to control sulfur accumulation in the sediments. However, in cores JB-04 and JB-05, some OC-rich sediments (containing more than 0.5% OC) do not accumulate sulfur ([Figure 4](#)). In these two cores, OC-rich, nonsulfidic layers are found above 3.5 m bgs, the average depth of the water table ([Figure 5](#)).

Present within the capillary fringe, these OC-rich sediments are seasonally exposed to oxygen and thus do not retain reduced sulfur. Low sulfide contents are also observed in permanently saturated parts of the aquifer in JB-05 and JB-02. Sediment textures reported in [Table S2](#) show that organic-rich layers in JB-03 and JB-05 are coarser-grained than in the other cores. In JB-02, core logs indicate a sediment grain-size transition between sandy nonreduced sediments (3.8 m bgs) to silt/clay-sized reduced sediments (4.1 m bgs) which is in agreement with the sharp variation in the sulfur profile shown in [Figure 2](#). Size-fractionated elemental analysis also shows sulfur concentrations increasing with decreasing particle size in each of the analyzed samples ([Table 1](#)). These results support the previous conclusions that NRZs accumulate sulfide and uranium, based on their low permeability, higher specific surface areas, smaller particle sizes, and elevated clay and uranium contents.[\(4\)](#) In more permeable sediments, oxygen can be transported faster than it is consumed, and anoxic conditions are difficult to achieve. Thus, sulfur is not reduced to lower-solubility sulfide in coarser sediments.

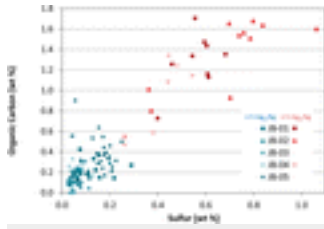


Figure 4. Elemental abundances of organic carbon and sulfur in the sediments from wells JB-01 (squares), JB-02 (triangles), JB-03 (circles), JB-04 (crosses), and JB-05 (diamonds). Red symbols: ≥ 5 mg/kg U; blue symbols: ≤ 5 mg/kg U.

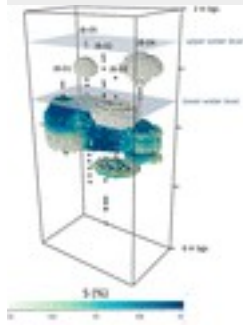


Figure 5. Three-dimensional representation of the distribution of organic-rich lenses within the JB well cluster. Delimitation of lenses is based on OC measurements ($OC > 0.4\%$) and macroscopic observations during sample collection. Within these lenses, sulfur concentration is interpolated on a regular grid from S measurements in core samples. White and colored volumes are the nonreduced zones and S-enriched NRZs, respectively.

Sulfur Biogeochemical Cycling

Sulfur K-edge XANES measurements of NRZ sediments showed the presence of elemental sulfur and iron monosulfides (mackinawite and mostly greigite) and disulfide (pyrite). Mackinawite, together with elemental sulfur, can be generated from sulfate by microbial activity, as well as by abiotic reduction of Fe(III) minerals coupled to the oxidation of aqueous sulfide. (3, 12, 13) The presence of these species in NRZ sediments strongly suggests ongoing sulfate reduction in these zones and implies an active mass exchange between the aquifer and the NRZ. Results also show the presence of a significant amount of mackinawite in aquifer sediments in the capillary fringe above the NRZ (Figure S4), indicating that reducing conditions are present in the oxic portion of the aquifer, probably maintained in the interior of aggregates within the organic-rich sediments.

Under low-oxygen conditions, two-thirds of Fe(II) in mackinawite can oxidize, resulting in a solid-state transformation to greigite. (14) The presence of greigite in all NRZ samples, together with the presence of some sulfate in the majority (Figure 3), suggests that NRZs are regularly exposed to oxidants that penetrate through these sediment structures. Elemental sulfur can also be produced

from abiotic or biotic FeS oxidation in natural sediments.[\(15\)](#) These observations point to dynamic redox cycling in the aquifer, likely tied to seasonal variations in groundwater elevation that enable sulfate reduction and FeS precipitation during periods of complete saturation and/or low oxygen conditions, and FeS oxidation to Fe₃S₄ and sulfate during partially saturated and/or higher oxygen conditions when groundwater elevations seasonally decrease.[\(16\)](#)

The composition of the reduced pool of sulfur varies with depth: near-surface oxidized sediments mostly contain mackinawite whereas pyrite is most abundant in the interior and lower portions of the NRZ ([Figure S5](#)). Greigite is most abundant at the NRZ–aquifer interface. This could suggest that greigite is a precursor for pyrite formation from mackinawite, as widely proposed.[\(10\)](#) However, this traditional view is strongly debated, with others suggesting that greigite and pyrite may represent distinct end points of competing Fe–S mineralization pathways.[\(14, 17\)](#) In the latter case, differences in the end-point mineralogy may be due to local differences in biomass activity and pore-water Fe²⁺ concentration.[\(17\)](#)

It is known that following sulfate reduction in sediments, aqueous sulfide can undergo reoxidation and react either with organic matter to form organic S species or with reactive ferric species to form iron sulfides.[\(11, 18, 19\)](#) Our results show that in Rifle NRZs sulfur pairs with iron and forms FeS particles in the fine fraction, whereas OC is most abundant in larger size fractions and no organic S species are detected with XANES analysis. Sulfurization of organic matter is not observed in these sediments to a detectable extent, and instead, the dominant sink for reduced sulfur seems to be iron sulfides. As suggested by Couture et al. (2013), this may be an indicator of the low lability of the OC.[\(19\)](#)

Biogeochemical Controls on Uranium

[Figure 4](#) shows that elevated uranium concentrations (red data points) are only found in sediments rich in OC: sediments with more than 5 mg_U/kg contain at least 0.2% S and 0.3% OC. However, OC-rich zones, when not enriched in sulfur (such as occurs in JB-05 core), do not accumulate uranium ([Figure S10](#)). This suggests that sulfate-reducing conditions are important for uranium accumulation. These sulfate-reducing conditions were established with the deposition of organic-rich sediments. The railroad was built on the floodplain in its current configuration by 1910. Assuming that the water table was then approximately in its present-day location, the NRZ sediments should have been buried and saturated as of that time, at least 40 years prior to arrival of uranium contamination. Sulfate-reducing conditions should have developed quickly. Consequently, the system is expected to have been sulfidic when dissolved U(VI) began to occur as a contaminant.

In reduced sediments, uranium is relatively more abundant in the largest size fraction (4.2 and 4.3 m) and smallest size fraction (4.4 m) as compared to the intermediate size. Sulfide is most abundant

in the smallest size fraction, whereas organic carbon is most abundant in the largest fraction. These results suggest that uranium is associated with both sulfides and with organic carbon ([Table 1](#)). These observations suggest that biological activity associated with decomposition of organic matter is important and it is plausible that U(VI) reduction is facilitated by binding to bacterial biomass, as previously proposed.[\(20\)](#) It is also plausible that electron shuttles are more abundant in these organic-enhanced zones and facilitate uranium reduction. Several reduction pathways are likely to be present in these sediments: abiotic reduction by iron sulfides that produce uraninite[\(21\)](#) as well as biological mediation of electron transfer via enzymatic activity.[\(20\)](#)

The abundant sulfides in NRZs are expected to react strongly with dissolved oxygen, protecting uranium from oxidation.[\(4\)](#) Other oxidants that are less reactive with sulfide, such as nitrate, may be more important for uranium oxidation in NRZs, as demonstrated by several authors.[\(22-25\)](#) Nitrate is highly soluble and can diffuse into NRZs to be subsequently reduced to nitrite, a strong oxidant for U(IV),[\(24-26\)](#) by denitrifying bacteria. Williams et al. (2014) have observed nitrate in relatively high (millimolar) concentrations in Rifle porewaters proximal to NRZ sediments during seasonal water level changes[\(27\)](#) and recently Nolan and Weber showed a correlation between uranium and nitrate levels in groundwater in major U.S. aquifers.[\(28\)](#) Further research is required to investigate the possible roles of denitrification and nitrification on uranium accumulation and release.

Depositional Environment

Recent papers have noted that deep burial of soil horizons is common in alluvial floodplain sediments[\(29, 30\)](#) and contribute large stores of carbon and biological productivity to deep (>30 cm) floodplain sediments.[\(31, 32\)](#) If Rifle NRZs are buried soil horizons, then they should share textural characteristics with modern nearby soil horizons. Sediment textures in local river banks were examined to evaluate this hypothesis. Bank sediments from the Roaring fork River, a nearby tributary of the Colorado River, exhibited strong textural similarity to NRZs ([Figure 1](#)). The riverbed was paved with cobbles of 50 to >200 mm diameter, common in the upper CRB.[\(33, 34\)](#) Riverbank grass mats growing over this cobble surface produced direct contact of fine-grained sediments and roots and underlying cobbles. A sulfidic horizontal redox transition located approximately at the water level ([Figure 1](#)), produced sharp juxtaposition of lighter colored and black sulfidic sediments. The presence of these textural features in Rifle NRZ sediments supports the idea that NRZs are buried riverbank soil horizons. It is reasonable to posit that NRZs are present in alluvial floodplains throughout the upper CRB and accumulate uranium wherever it is present in groundwater.

Environmental Implications

Storage and release of uranium, OC, and sulfur in NRZs may have a substantial impact on aquifer water quality. NRZ sediments analyzed in this study contain an average of 26 mg_U/kg. Based on the

inferred dimensions of the main NRZ observed here (4 m long, 2 m wide and about 0.5 m thick), and using a sediment particle density of 2.75 kg/L,(35) the estimated inventory of uranium within the limited area of the study plot is 286 g. The net flux of uranium from the NRZ to the aquifer determines whether this NRZ could sustain a uranium plume. Assuming a steady-state flow system with a fixed flux independent of the concentrations and reaction rates within the NRZ, this flux J_{NRZ} is given by $J_{\text{NRZ}} = C \frac{\phi h v}{x}$ where ϕ is the porosity (0.15–0.35(36)), h is the saturated thickness of the aquifer (2.4–3.4 m(36)), v is the mean pore water velocity by (109–182 m/yr(36)), and x is the downgradient flow path (138 m from the NRZ to the river). The concentration of U in the groundwater (C) is the threshold value used to define a contaminant plume at a given value of x . Using this formulation, the flux of U required to maintain plume levels above the UMTRA groundwater remediation standard for uranium of 44 $\mu\text{g}/\text{kg}$ (6) is $J_{\text{NRZ}} = 13\text{--}69 \text{ mg}/\text{m}^2/\text{yr}$. The time interval over which a plume could be maintained, for a simple order-of-magnitude estimate, depends on the inventory of uranium with the NRZ, the interfacial area of the NRZs and the net flux. Assuming 286 g of uranium in the NRZ released to the aquifer through its top and bottom interfaces (a 16 m^2 area), then the maximum duration over which a 2m-wide plume could be sustained is 259–1375 years. Measuring the magnitude of the uranium outward flux from NRZs is difficult and beyond the scope of the present study. Nevertheless, this discussion highlights the need to understand the rates and mechanisms of uranium mass transfer between NRZs and the aquifer under field conditions, the processes that govern the generation of and transport of U within the NRZ, as well as the potential importance of NRZs to plume generation when oxidizing conditions are present. Large climate-driven variations in redox conditions occur seasonally and interseasonally (e.g., multiyear drought cycles).(16) Consequently, oxidation of NRZs is likely to be common, and decreasing rainfall in future seasons would be expected to overall drive uranium release through enhanced oxidation of uranium in the unsaturated zone accompanied by seasonal flushing of U during meltwater discharge events. Because organic-enriched sediments are expected to be common floodplain depositional features, we should expect them to be present in the numerous contaminated floodplain DOE sites in the upper CRB.

Supporting Information

The Supporting Information is available free of charge on the [ACS Publications website](https://pubs.acs.org) at DOI: [10.1021/acs.est.5b03208](https://doi.org/10.1021/acs.est.5b03208).

- Additional method details, figures, and tables ([PDF](#))
- **PDF**
 - [es5b03208_si_001.pdf \(4.08 MB\)](#)

**Physico-Chemical Heterogeneity of Organic-Rich Sediments in the Rifle Aquifer, CO:
Impact on Uranium Biogeochemistry**

Supplemental Information for:

Physico-chemical heterogeneity of organic-rich sediments in the Rifle aquifer, CO: Impact on uranium biogeochemistry

*Noémie Janot^a, Juan S. Lezama Pacheco^{a,b}, Don Q. Pham^a, Timothy M. O'Brien^a,
Debra Hausladen^b, Vincent Noël^a, Florent Lallier^c, Kate Maher^d, Scott Fendorf^b, Kenneth H.
Williams^e, Philip E. Long^e, John R. Bargar^{a,*}*

^a Stanford Synchrotron Radiation Lightsource, SLAC National Accelerator Laboratory, Menlo
Park, CA 94025, United States

^b Department of Environmental Earth System Science, Stanford University, Stanford, CA 94305,
United States

^c Université de Lorraine, Vandœuvre-lès-Nancy, F-54501, France

^d Department of Geological Sciences, Stanford University, Stanford, CA 94305, United States

^d Earth Sciences Division, Lawrence Berkeley National Laboratory, Berkeley, CA 94720, United
States

Share [Download](#)

#Author Present Address

(N.J.) Université de Lorraine, LIEC, UMR7360, Vandœuvre-lès-Nancy, F-54501, France.

∇Author Present Address

(T.M.O.) Department of Geological Science, Stanford University, Stanford, CA 94305, United States.

◆Author Present Address

(F.L.) TOTAL E&P UK Ltd., Altens, Aberdeen AB12 3FG, United Kingdom.

The authors declare no competing financial interest.

•

Acknowledgment

The SLAC Scientific Focus Area (SFA) program is supported by U.S. DOE Office of Biological and Environmental Research (BER), Subsurface Biogeochemical Research (SBR) program under subcontract DE-AC02-76SF00515. Logistical support was provided by Rifle field research program at the Lawrence Berkeley National Laboratory through funding from DOE-BER-SBR to the Sustainable Systems SFA 2.0 under contract DE-AC02-05CH11231. We thank Julian Carpenter and Dr. Ritimukta Sarangi for providing FeS and elemental sulfur reference compounds, respectively. Greigite synthesis was developed at the IMPMC (Paris, France).

• [Reference QuickView](#)

•

References

This article references 36 other publications.

1. [1.](#)

Davis, J. A.; Curtis, G. P.; Wilkins, M. J.; Kohler, M.; Fox, P. M.; Naftz, D. L.; Lloyd, J. R. Processes affecting transport of uranium in a suboxic aquifer *Phys. Chem. Earth, Parts A/B/C* **2006**, 31 (10–14) 548– 555 DOI: 10.1016/j.pce.2006.04.005

[\[Crossref\]](#)

2. [2.](#)

Zachara, J. M.; Long, P. E.; Bargar, J. R.; Davis, J. A.; Fox, P. M.; Fredrickson, J. K.; Freshley, M. D.; Konopka, A. E.; Liu, C.; McKinley, J. P.; Rockhold, M. L.; Williams, K. H.; Yabusaki, S. B. Persistence of uranium groundwater plumes: Contrasting mechanisms at two DOE sites in the groundwater-river interaction zone *J. Contam. Hydrol.* **2013**, 147, 45– 72 DOI: 10.1016/j.jconhyd.2013.02.001

[\[Crossref\]](#), [\[PubMed\]](#), [\[CAS\]](#)

3. [3.](#)

Williams, K. H.; Long, P. E.; Davis, J. A.; Wilkins, M. J.; N'Guessan, A. L.; Steefel, C. I.; Yang, L.; Newcomer, D. R.; Spane, F. A.; Kerkhof, L. J.; McGuinness, L.; Dayvault, R.; Lovley, D. R. Acetate availability and its influence on sustainable bioremediation of uranium-contaminated groundwater *Geomicrobiol. J.* **2011**, 28 (5–6) 519– 539 DOI: 10.1080/01490451.2010.520074

[\[Crossref\]](#), [\[CAS\]](#)

4. [4.](#)

Campbell, K. M.; Kukkadapu, R. K.; Qafoku, N. P.; Peacock, A. D.; Leshner, E.; Williams, K. H.; Bargar, J. R.; Wilkins, M. J.; Figueroa, L.; Ranville, J. F.; Davis, J. A.; Long, P. E. Geochemical, mineralogical and microbiological characteristics of sediment from a naturally reduced zone in a uranium-contaminated aquifer *Appl. Geochem.* **2012**, 27 (8) 1499– 1511 DOI: 10.1016/j.apgeochem.2012.04.013

[\[Crossref\]](#), [\[CAS\]](#)

5. [5.](#)

Qafoku, N. P.; Gartman, B. N.; Kukkadapu, R. K.; Arey, B. W.; Williams, K. H.; Mouser, P. J.; Heald, S. M.; Bargar, J. R.; Janot, N.; Yabusaki, S.; Long, P. E. Geochemical and mineralogical investigation of uranium in multi-element contaminated, organic-rich subsurface sediment *Appl. Geochem.* **2014**, 42, 77– 85 DOI: 10.1016/j.apgeochem.2013.12.001

[\[Crossref\]](#), [\[CAS\]](#)

6. [6.](#)

U.S. DOE Grand Junction Office. Final Site Observational Work Plan for the UMTRA Project Old Rifle Site; **1999**.

7. [7.](#)

Anderson, R. T.; Vrionis, H. A.; Ortiz-bernad, I.; Resch, C. T.; Long, P. E.; Dayvault, R. D.; Karp, K.; Marutzky, S.; Metzler, D. R.; Peacock, A. D.; White, D. C.; Lowe, M.; Lovley, D. R. Stimulating the in situ activity of *Geobacter* species to remove uranium from the

groundwater of a uranium-contaminated aquifer *Appl. Environ. Microbiol.* **2003**, 69 (10) 5884– 5891 DOI: 10.1128/AEM.69.10.5884-5891.2003

[\[Crossref\]](#), [\[PubMed\]](#), [\[CAS\]](#)

8. [8.](#)

Yabusaki, S. B.; Fang, Y.; Long, P. E.; Resch, C. T.; Peacock, A. D.; Komlos, J.; Jaffé, P. R.; Morrison, S. J.; Dayvault, R. D.; White, D. C.; Anderson, R. T. Uranium removal from groundwater via in situ biostimulation: Field-scale modeling of transport and biological processes *J. Contam. Hydrol.* **2007**, 93 (1–4) 216– 235 DOI: 10.1016/j.jconhyd.2007.02.005

[\[Crossref\]](#), [\[PubMed\]](#), [\[CAS\]](#)

9. [9.](#)

U.S. DOE Legacy Management. Rifle, Colorado, Processing Sites and Disposal Site - Fact Sheet; **2011**.

10. [10.](#)

Qafoku, N. P.; Kukkadapu, R. K.; McKinley, J. P.; Arey, B. W.; Kelly, S. D.; Wang, C.; Resch, C. T.; Long, P. E. Uranium in framboidal pyrite from a naturally bioreduced alluvial sediment *Environ. Sci. Technol.* **2009**, 43 (22) 8528– 8534 DOI: 10.1021/es9017333

[\[ACS Full Text\]](#) , [\[CAS\]](#)

11. [11.](#)

Prietzl, J.; Thieme, J.; Tyufekchieva, N.; Paterson, D.; McNulty, I.; Kögel-Knabner, I. Sulfur speciation in well-aerated and wetland soils in a forested catchment assessed by sulfur K-edge X-ray absorption near-edge spectroscopy (XANES) *J. Plant Nutr. Soil Sci.* **2009**, 172 (3) 393– 403 DOI: 10.1002/jpln.200800054

[\[Crossref\]](#), [\[CAS\]](#)

12. [12.](#)

Druhan, J. L.; Steefel, C. I.; Conrad, M. E.; DePaolo, D. J. A large column analog experiment of stable isotope variations during reactive transport: I. A comprehensive model of sulfur cycling and $\delta^{34}\text{S}$ fractionation *Geochim. Cosmochim. Acta* **2014**, 124, 366– 393 DOI: 10.1016/j.gca.2013.08.037

[\[Crossref\]](#), [\[CAS\]](#)

13. [13.](#)

Wei, D.; Osseo-Asare, K. Aqueous synthesis of finely divided pyrite particles *Colloids Surf., A* **1997**, 121 (1)27– 36 DOI: 10.1016/S0927-7757(95)03502-8

[\[Crossref\]](#), [\[CAS\]](#)

14. [14.](#)

Rickard, D. T.; Luther, G. W. Chemistry of iron sulfides *Chem. Rev.* **2007**, 107 (2) 514– 562 DOI: 10.1021/cr0503658

[\[ACS Full Text\]](#), [\[CAS\]](#)

15. [15.](#)

Burton, E. D.; Bush, R. T.; Sullivan, L. A.; Hocking, R. K.; Mitchell, D. R. G.; Johnston, S. G.; Fitzpatrick, R. W.; Raven, M. D.; McClure, S.; Jang, L. Y. Iron-monosulfide oxidation in natural sediments: resolving microbially mediated S transformations using XANES, electron microscopy, and selective extractions *Environ. Sci. Technol.* **2009**, 43 (9) 3128– 3134 DOI: 10.1021/es8036548

[\[ACS Full Text\]](#), [\[CAS\]](#)

16. [16.](#)

Lezama-Pacheco, J. S.; Cerrato, J. M.; Veeramani, H.; Alessi, D. S.; Suvorova, E.; Bernier-Latmani, R.; Giammar, D. E.; Long, P. E.; Williams, K. H.; Bargar, J. R. Long-term in situ oxidation of biogenic uraninite in an alluvial aquifer: Impact of dissolved oxygen and calcium *Environ. Sci. Technol.* **2015**, 49 (12) 7340– 7347 DOI: 10.1021/acs.est.5b00949

[\[ACS Full Text\]](#), [\[CAS\]](#)

17. [17.](#)

Burton, E. D.; Bush, R. T.; Johnston, S. G.; Sullivan, L. A.; Keene, A. F. Sulfur biogeochemical cycling and novel Fe–S mineralization pathways in a tidally re-flooded wetland *Geochim. Cosmochim. Acta* **2011**, 75(12) 3434– 3451 DOI: 10.1016/j.gca.2011.03.020

[\[Crossref\]](#), [\[CAS\]](#)

18. [18.](#)

Zeng, T.; Arnold, W. A.; Toner, B. M. Microscale characterization of sulfur speciation in lake sediments *Environ. Sci. Technol.* **2013**, 47 (3) 1287– 1296 DOI: 10.1021/es303914q

[\[ACS Full Text\]](#), [\[CAS\]](#)

19. [19.](#)

Couture, R.-M.; Wallschläger, D.; Rose, J.; Van Cappellen, P. Arsenic binding to organic and inorganic sulfur species during microbial sulfate reduction: a sediment flow-through reactor experiment *Environ. Chem.* **2013**, 10 (4) 285 DOI: 10.1071/EN13010

[\[Crossref\]](#), [\[CAS\]](#)

20. [20.](#)

Bargar, J. R.; Williams, K. H.; Campbell, K. M.; Long, P. E.; Stubbs, J. E.; Suvorova, E. I.; Lezama-Pacheco, J. S.; Alessi, D. S.; Stylo, M.; Webb, S. M.; Davis, J. A.; Giammar, D. E.; Blue, L. Y.; Bernier-Latmani, R. Uranium redox transition pathways in acetate-amended sediments *Proc. Natl. Acad. Sci. U. S. A.* **2013**, 110 (12) 4506– 4511 DOI: 10.1073/pnas.1219198110

[\[Crossref\]](#), [\[CAS\]](#)

21. [21.](#)

Hyun, S. P.; Davis, J. a.; Sun, K.; Hayes, K. F. Uranium(VI) reduction by iron(II) monosulfide mackinawite *Environ. Sci. Technol.* **2012**, 46 (6) 3369– 3376 DOI: 10.1021/es203786p

[\[ACS Full Text\]](#) , [\[CAS\]](#)

22. [22.](#)

Senko, J. M.; Istok, J. D.; Suflita, J. M.; Krumholz, L. R. In-situ evidence for uranium immobilization and remobilization *Environ. Sci. Technol.* **2002**, 36 (7) 1491– 1496 DOI: 10.1021/es011240x

[\[ACS Full Text\]](#) , [\[CAS\]](#)

23. [23.](#)

Istok, J. D.; Senko, J. M.; Krumholz, L. R.; Watson, D.; Bogle, M. A.; Peacock, A.; Chang, Y.-J.; White, D. C. In situ bioreduction of technetium and uranium in a nitrate-contaminated aquifer *Environ. Sci. Technol.* **2004**, 38 (2) 468– 475 DOI: 10.1021/es034639p

[\[ACS Full Text\]](#) , [\[CAS\]](#)

24. [24.](#)

Moon, H. S.; Komlos, J.; Jaffé, P. R. Uranium reoxidation in previously bioreduced sediment by dissolved oxygen and nitrate *Environ. Sci. Technol.* **2007**, 41 (13) 4587– 4592 DOI: 10.1021/es063063b

[\[ACS Full Text\]](#) , [\[CAS\]](#)

25. [25.](#)

Moon, H. S.; Komlos, J.; Jaffé, P. R. Biogenic U(IV) oxidation by dissolved oxygen and nitrate in sediment after prolonged U(VI)/Fe(III)/SO₄(²⁻) reduction *J. Contam. Hydrol.* **2009**, 105 (1– 2) 18– 27 DOI: 10.1016/j.jconhyd.2008.10.014

[\[Crossref\]](#), [\[PubMed\]](#), [\[CAS\]](#)

26. [26.](#)

Carpenter, J.; Bi, Y.; Hayes, K. F. Influence of iron sulfides on abiotic oxidation of UO₂ by nitrite and dissolved oxygen in natural sediments *Environ. Sci. Technol.* **2015**, 49, 1078– 1085 DOI: 10.1021/es504481n

[\[ACS Full Text\]](#), [\[CAS\]](#)

27. [27.](#)

Williams, K. H.; Bargar, J. R. A Floodplain perspective on elemental cycling and uranium plume persistence. In Goldschmidt Conference, **2014**.

28. [28.](#)

Nolan, J.; Weber, K. A. Natural uranium contamination in major U.S. aquifers linked to nitrate *Environ. Sci. Technol. Lett.* **2015**, 2 (8) 215– 220 DOI: 10.1021/acs.estlett.5b00174

[\[ACS Full Text\]](#), [\[CAS\]](#)

29. [29.](#)

Gurwick, N. P.; Groffman, P. M.; Yavitt, J. B.; Gold, A. J.; Blazejewski, G. A.; Stolt, M. H. Microbially available carbon in buried riparian soils in a glaciated landscape *Soil Biol. Biochem.* **2008**, 40 (1) 85– 96 DOI: 10.1016/j.soilbio.2007.07.007

[\[Crossref\]](#), [\[CAS\]](#)

30. [30.](#)

Blazejewski, G. A.; Stolt, M. H.; Gold, A. J.; Gurwick, N. P.; Groffman, P. M. Spatial distribution of carbon in the subsurface of riparian zones *Soil Sci. Soc. Am. J.* **2009**, 73 (5) 1733 DOI: 10.2136/sssaj2007.0386

[\[Crossref\]](#), [\[CAS\]](#)

31. [31.](#)

Ricker, M. C.; Stolt, M. H.; Donohue, S. W.; Blazejewski, G. A.; Zavada, M. S. Soil organic carbon pools in riparian landscapes of southern New England *Soil Sci. Soc. Am. J.* **2013**, 77 (3) 1070 DOI: 10.2136/sssaj2012.0297

[\[Crossref\]](#), [\[CAS\]](#)

32. [32.](#)

Chaopricha, N. T.; Marín-Spiotta, E. Soil burial contributes to deep soil organic carbon storage *Soil Biol. Biochem.* **2014**, 69, 251– 264 DOI: 10.1016/j.soilbio.2013.11.011

[\[Crossref\]](#), [\[CAS\]](#)

33. [33.](#)

Elliott, J. G.; Hammack, L. A. Entrainment of riparian gravel and cobbles in an alluvial reach of a regulated canyon river *River Res. Appl.* **2000**, 16 (1) 37– 50 DOI: 10.1002/(SICI)1099-1646(200001/02)16:1<37::AID-RRR564>3.0.CO;2-V

[\[Crossref\]](#)

34. [34.](#)

Van Steeter, M. M.; Pitlick, J. Geomorphology and endangered fish habitats of the upper Colorado River: 1. Historic changes in streamflow, sediment load, and channel morphology *Water Resour. Res.* **1998**, 34 (2)287 DOI: 10.1029/97WR02766

[\[Crossref\]](#)

35. [35.](#)

Fang, Y.; Yabusaki, S. B.; Morrison, S. J.; Amonette, J. P.; Long, P. E. Multicomponent reactive transport modeling of uranium bioremediation field experiments *Geochim. Cosmochim. Acta* **2009**, 73 (20) 6029–6051 DOI: 10.1016/j.gca.2009.07.019

[\[Crossref\]](#), [\[CAS\]](#)

36. [36.](#)

Yabusaki, S. B.; Fang, Y.; Williams, K. H.; Murray, C. J.; Ward, A. L.; Dayvault, R. D.; Waichler, S. R.; Newcomer, D. R.; Spane, F. A.; Long, P. E. Variably saturated flow and multicomponent biogeochemical reactive transport modeling of a uranium bioremediation field experiment *J. Contam. Hydrol.* **2011**, 126 (3–4) 271– 290 DOI: 10.1016/j.jconhyd.2011.09.002

[\[Crossref\]](#), [\[PubMed\]](#), [\[CAS\]](#)

<ul style="list-style-type: none">Citing Articles	
<ul style="list-style-type: none">Related Content	

Citation data is made available by participants in [Crossref's](#) Cited-by Linking service. For a more comprehensive list of citations to this article, users are encouraged to perform a search in [SciFinder](#).

•

[Carbonate Facilitated Mobilization of Uranium from Lacustrine Sediments under Anoxic Conditions](#)

Marina Seder-ColominaArnaud MangeretLucie StettenPauline MerrotOlivier DiezAnthony JulienEvelyne BarkerAntoine ThouvenotJohn BargarCharlotte CazalaGuillaume Morin

Environmental Science & Technology **2018** Article ASAP

[Abstract](#) | [Full Text HTML](#) | [PDF](#) | [PDF w/ Links](#)

•

[Uranium Retention in a Bioreduced Region of an Alluvial Aquifer Induced by the Influx of Dissolved Oxygen](#)

Donald PanKenneth H. WilliamsMark J. RobbinsKarrie A. Weber

Environmental Science & Technology **2018** 52 (15), 8133-8145

[Abstract](#) | [Full Text HTML](#) | [PDF](#) | [PDF w/ Links](#)

•

[Oxidative Uranium Release from Anoxic Sediments under Diffusion-Limited Conditions](#)

Sharon E. Bone, Melanie R. Cahill, Morris E. Jones, Scott Fendorf, James Davis, Kenneth H. Williams, and John R. Bargar

Environmental Science & Technology **2017** 51 (19), 11039-11047

[Abstract](#) | [Full Text HTML](#) | [PDF](#) | [PDF w/ Links](#)

•

[Redox Controls over the Stability of U\(IV\) in Floodplains of the Upper Colorado River Basin](#)

Vincent Noël, Kristin Boye, Juan S. Lezama Pacheco, Sharon E. Bone, Noémie Janot, Emily Cardarelli, Kenneth H. Williams, and John R. Bargar
Environmental Science & Technology **2017** 51 (19), 10954-10964

[Abstract](#) | [Full Text HTML](#) | [PDF](#) | [PDF w/ Links](#)

•

[Production of Hydrogen Peroxide in Groundwater at Rifle, Colorado](#)

Xiu Yuan, Peter S. Nico, Xiang Huang, Tongxu Liu, Craig Ulrich, Kenneth H. Williams, and James A. Davis
Environmental Science & Technology **2017** 51 (14), 7881-7891

[Abstract](#) | [Full Text HTML](#) | [PDF](#) | [PDF w/ Links](#)

•

[Water Table Dynamics and Biogeochemical Cycling in a Shallow, Variably-Saturated Floodplain](#)

Steven B. Yabusaki, Michael J. Wilkins, Yilin Fang, Kenneth H. Williams, Bhavna Arora, John Bargar, Harry R. Beller, Nicholas J. Bouskill, Eoin L. Brodie, John N. Christensen, Mark E. Conrad, Robert E. Danczak, Eric King, Mohamad R. Soltanian, Nicolas F. Spycher, Carl I. Steefel, Tetsu K. Tokunaga, Roelof Versteeg, Scott R. Waichler, and Haruko M. Wainwright

Environmental Science & Technology **2017** 51 (6), 3307-3317

[Abstract](#) | [Full Text HTML](#) | [PDF](#) | [PDF w/ Links](#)

•

[Uranium Isotopic Fractionation Induced by U\(VI\) Adsorption onto Common Aquifer Minerals](#)

N. E. Jemison, T. M. Johnson, A. E. Shiel, and C. C. Lundstrom
Environmental Science & Technology **2016** 50 (22), 12232-12240

[Abstract](#) | [Full Text HTML](#) | [PDF](#) | [PDF w/ Links](#)

•

[Using strontium isotopes to evaluate the spatial variation of groundwater recharge](#)

John N. Christensen, Baptiste Dafflon, Alyssa E. Shiel, Tetsu K. Tokunaga, Jiamin Wan, Boris Faybishenko, Wenming Dong, Kenneth H. Williams, Chad Hobson, Shaun T. Brown, Susan S. Hubbard
Science of The Total Environment **2018** 637-638, 672-685

•

[Sulfate reduction processes in salt marshes affected by phosphogypsum: Geochemical influences on contaminant mobility](#)

Rafael Pérez-López, Sergio Carrero, Pablo Cruz-Hernández, Maria P. Asta, Francisco Macías, Carlos R. Cánovas, Clara Guglieri, José Miguel Nieto
Journal of Hazardous Materials **2018** 350, 154-161

•

[Uranium in groundwater - A synopsis based on a large hydrogeochemical data set](#)

Thomas Riedel, Christine Kübeck
Water Research **2018** 129, 29-38

•

[Hot Spots and Hot Moments of Nitrogen in a Riparian Corridor](#)

Dipankar Dwivedi, Bhavna Arora, Carl I. Steefel, Baptiste Dafflon, Roelof Versteeg
Water Resources Research **2018** 54 (1), 205-222

•

[Sulfidation mechanisms of Fe\(III\)-\(oxyhydr\)oxide nanoparticles: a spectroscopic study](#)

Naresh Kumar, Juan Lezama Pacheco, Vincent Noël, Gabrielle Dublet, Gordon E. Brown

- Environmental Science: Nano* **2018** 5 (4), 1012-1026
- [**Understanding controls on redox processes in floodplain sediments of the Upper Colorado River Basin**](#)
 Vincent Noël, Kristin Boye, Ravi K. Kukkadapu, Sharon Bone, Juan S. Lezama Pacheco, Emily Cardarelli, Noémie Janot, Scott Fendorf, Kenneth H. Williams, John R. Bargar
Science of The Total Environment **2017** 603-604, 663-675
 - [**Characterization of natural organic matter in low-carbon sediments: Extraction and analytical approaches**](#)
 Patricia M. Fox, Peter S. Nico, Malak M. Tfaily, Katherine Heckman, James A. Davis
Organic Geochemistry **2017** 114, 12-22
 - [**Partitioning of uranyl between ferrihydrite and humic substances at acidic and circum-neutral pH**](#)
 Gabrielle Dublet, Juan Lezama Pacheco, John R. Bargar, Scott Fendorf, Naresh Kumar, Gregory V. Lowry, Gordon E. Brown
Geochimica et Cosmochimica Acta **2017** 215, 122-140
 - [**Abundance and Distribution of Microbial Cells and Viruses in an Alluvial Aquifer**](#)
 Donald Pan, Jason Nolan, Kenneth H. Williams, Mark J. Robbins, Karrie A. Weber
Frontiers in Microbiology **2017** 8,
 - [**Transport and humification of dissolved organic matter within a semi-arid floodplain**](#)
 Wenming Dong, Jiamin Wan, Tetsu K. Tokunaga, Benjamin Gilbert, Kenneth H. Williams
Journal of Environmental Sciences **2017** 57, 24-32
 - [**Thermodynamically controlled preservation of organic carbon in floodplains**](#)
 Kristin Boye, Vincent Noël, Malak M. Tfaily, Sharon E. Bone, Kenneth H. Williams, John R. Bargar, Scott Fendorf
Nature Geoscience **2017** 10 (6), 415-419
 - [**Uranium delivery and uptake in a montane wetland, north-central Colorado, USA**](#)
 R. Randall Schumann, Robert A. Zielinski, James K. Otton, Michael P. Pantea, William H. Orem
Applied Geochemistry **2017** 78, 363-379
 - [**Expanding the role of reactive transport models in critical zone processes**](#)
 Li Li, Kate Maher, Alexis Navarre-Sitchler, Jenny Druhan, Christof Meile, Corey Lawrence, Joel Moore, Julia Perdril, Pamela Sullivan, Aaron Thompson, Lixin Jin, Edward W. Bolton, Susan L. Brantley, William E. Dietrich, K. Ulrich Mayer, Carl I. Steefel, Albert Valocchi, John Zachara, Benjamin Kocar, Jennifer McIntosh, Benjamin M. Tutolo, Mukesh Kumar, Eric Sonnenthal, Chen Bao, Joe Beisman
Earth-Science Reviews **2017** 165, 280-301
 - [**Metatranscriptomic Analysis Reveals Unexpectedly Diverse Microbial Metabolism in a Biogeochemical Hot Spot in an Alluvial Aquifer**](#)
 Talia N. M. Jewell, Ulas Karaoz, Markus Bill, Romy Chakraborty, Eoin L. Brodie, Kenneth H. Williams, Harry R. Beller
Frontiers in Microbiology **2017** 8,
 - [**Uranium\(IV\) adsorption by natural organic matter in anoxic sediments**](#)
 Sharon E. Bone, James J. Dynes, John Cliff, John R. Bargar
Proceedings of the National Academy of Sciences **2017** 114 (4), 711-716
 - [**Uranium mobility and accumulation along the Rio Paguete, Jackpile Mine in Laguna Pueblo, NM**](#)
 Johanna M. Blake, Cherie L. De Vore, Sumant Avasarala, Abdul-Mehdi Ali, Claudia Roldan, Fenton Bowers, Michael N. Spilde, Kateryna Artyushkova, Matthew F. Kirk, Eric Peterson, Lucia Rodriguez-Freire, José M. Cerrato
Environmental Science: Processes & Impacts **2017** 19 (4), 605-621
 - [**Identifying geochemical hot moments and their controls on a contaminated river floodplain system using wavelet and entropy approaches**](#)
 Bhavna Arora, Dipankar Dwivedi, Susan S. Hubbard, Carl I. Steefel, Kenneth H. Williams

Environmental Modelling & Software **2016** 85, 27-41

- [**Hierarchical Bayesian method for mapping biogeochemical hot spots using induced polarization imaging**](#)

Haruko M. Wainwright, Adrian Flores Orozco, Matthias B?cker, Baptiste Dafflon, Jinsong Chen, Susan S. Hubbard, Kenneth H. Williams

Water Resources Research **2016** 52 (1), 533-551