Cross-linked cyclodextrin-based material for treatment of metals and organic substances present in industrial discharge waters

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To cite this version:


HAL Id: hal-02636870
https://hal.inrae.fr/hal-02636870
Submitted on 27 May 2020

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Abstract

In this study, a polymer, prepared by crosslinking cyclodextrin (CD) by means of a polycarboxylic acid, was used for the removal of pollutants from spiked solutions and discharge waters from the surface treatment industry. In spiked solutions containing five metals, sixteen polycyclic aromatic hydrocarbons (PAH) and three alkylphenols (AP), the material exhibited high adsorption capacities: >99% of Co\(^{2+}\), Ni\(^{2+}\) and Zn\(^{2+}\) were removed, between 65 and 82% of the PAHs, as well as 69 to 90% of the APs. Due to the structure of the polymer and its specific characteristics, such as the presence of carboxylic groups and CD cavities, the adsorption mechanism involves four main interactions: ion exchange, electrostatic interactions and precipitation for metal removal, and inclusion complexes for organics removal. In industrial discharge waters, competition effects appeared, especially because of the presence of calcium at high concentrations, which competed with other pollutants for the adsorption sites of the adsorbent.

Introduction

Although considerable efforts have been made by the industrial sector over the last 30 years, the problem of water pollution still remains a significant concern. Particularly affected by this issue are the discharge waters (DWs) of the surface treatment (ST) industries known for using large amounts of water and chemicals in their manufacturing processes. Despite these industries...
have their own treatment plants, generally physicochemical
decontamination steps, the DWs still contain non-negligible
amounts of pollutants. Among them, metals (in particular
chromium, nickel and zinc) are commonly found at concentra-
tions in the range of milligrams per liter, and organic molecules,
such as polycyclic aromatic hydrocarbons (PAHs) and
alkylphenols (APs) at concentrations varying from hundreds of
nanograms per liter to some micrograms per liter [1].

However, it is extremely difficult to remove pollutants present
at low concentrations (a few hundreds of micrograms per liter
for some organic substances in DWs). For this purpose, specific
systems can be added, called effluent finishing treatments. A
sequential dual approach can be considered: firstly, adsorption
onto carbon to remove organics (e.g., solvents, oils, PAHs and
volatile organic compounds) combined with ion-exchange and/
or chelation by means of organic resins to remove inorganic
pollutants (e.g., metals and anions such as fluorides). Charles et
al. [2] recently reported that this type of sequence is acknowl-
dged for its efficiency. However, it is an approach to water
treatment that combines two methods of separation using two
distinct commercial materials. To our knowledge, materials able
to combine the two functions are rare. Recently, bifunctional
natural derivatives have been proposed for this purpose. For
example, Zhao et al. [3] proposed a new cyclodextrin-based
material for the simultaneous adsorption of metals and cationic
dyes. Zhang et al. [4] studied the removal of cobalt and 1-naph-
thol onto magnetic nanoparticles containing cyclodextrin and
iron and Yang et al. [5] proposed a new nanocomposite adsor-
bent for the simultaneous removal of organic and inorganic
substances from water.

Here, we propose to use a single cross-linked cyclodextrin-
based polymer for the removal of metals and organic pollutants
present in polyc Contaminated effluents. Cyclodextrins (CDs),
synthetic substances obtained from the enzymatic degradation
of starch, belong to the family of cage molecules. They present
remarkable encapsulation properties leading to a host–guest
relationship with organic substances [6-9]. These cyclic oligo-
saccharides are water soluble in their native form and are often
modified to prepare novel insoluble CD-based materials. Two
patents published by Martel et al. [10], and Trotta et al. [11] can
be consulted for the use of carboxylic acids and pyromellitic
dianhydride, respectively, as agents to cross-linking CDs. Other
cross-linking agents such as epichlorohydrin, ethylene glycol
diglycidyl ether, glutaraldehyde, benzoquinone or isocyanates
can be also used [1,12].

The main objective of the study was to investigate the adsorp-
tion capacities of a non-conventional and versatile CD-based
material crosslinked with 1,2,3,4-butanetetracarboxylic acid
(BTCA) toward several inorganic and organic elements. Perform-
ances of such systems were evaluated in the presence of
spiked solutions and real DWs from ST industry containing five
metals, 16 PAHs, three APs (model pollutants present in DWs
from treatment-surface industries) in the presence and in the
absence of calcium. The polymer showed high adsorption
capacities in spiked solutions but adsorption strongly decreased
in discharge waters due to some competition effects, notably be-
tween inorganics for adsorption sites.

Results
Material characterization
The cross-linked polymer used in this study is presented in
Figure 1. In the control test (solution without pollutant),
polymer addition led to large pH variations. The non-activated
polyBTCA-CD (COOH form) decreased the pH value from 6 to
4.3 after 20 min of shaking whereas activated polyBTCA-CD
(COOO’Na+ form) led to a pH increase from 6 to 7.2 after 5 min
shaking. In both cases the final pH remained constant over
several hours.

The ion-exchange capacity (IEC) was equal to 0.705 mmol
of COOH functions per gram of polymer. The point of zero charge
(PZC) was plotted and followed the linear equations
\[
y = -0.9639x + 6.1422 \quad \text{and} \quad y = -0.9233x + 3.138
\]
indicating a pH of 6.4 and 3.4 for the PZC of activated and non-activated
polyBTCA-CD, respectively (Figure 2).

Figure 3 shows the XRD pattern of non-activated and activated
polymers. These diffractograms indicate a wide amorphous
peak between values of 20 of 10 and 30°, highlighting that poly-
mers are amorphous. No significant differences were observed
between the two types of polymer.

Solid-state $^{13}$C NMR spectra of the polymer are presented in
Figure 4. The cross polarization magic angle spinning
(CPMAS) spectrum shows the peaks of disordered cyclodextrin
(broad signals) in the range of 50–110 ppm. Three strong broad
bands attributable to the glucopyranose unit can be observed.
The peak at 101 ppm is attributed to the anomeric carbon C-1:
this confirms the presence of glucose units in the polymer. In
the range of 50–110 ppm, the CH$_2$ signals of CD (C-6, C-6’
and C-7) are completely hidden by the C-2, C-3 and C-5 peaks of
the glucopyranose units. In the MAS spectrum, three CH$_2$
signals are clear and the signals at 65, 62.1 and 60.2 ppm are
attributable to CH$_2$ in positions C-6’, C-6 and C-7, respectively.
These attributions were confirmed using cross-polarization with
a polarization inversion sequence (CPPI spectrum not shown),
which revealed three CH$_2$ signals due to the CD and also a peak
at 30.15 ppm corresponding to CH$_2$ groups introduced by the
cross-linking agent. As expected, the signals of the BTCA

$^{13}$C NMR spectra of the polymer are presented in
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a polarization inversion sequence (CPPI spectrum not shown),
which revealed three CH$_2$ signals due to the CD and also a peak
at 30.15 ppm corresponding to CH$_2$ groups introduced by the
cross-linking agent. As expected, the signals of the BTCA
crosslinking agent can be clearly distinguished (labeled a, b, c and d). In particular, the carbon of the carboxylic groups appears at 172.2 ppm. This peak (a,d) corresponds to esterified and free carboxylic groups of BTCA present in free carboxylic acids and in ester crosslinks, respectively. In the CPMAS spectrum we also note the presence of additional peaks due to the hydroxypropyl group present in the CD, and in particular the CH₃ group (C-9 carbon) at 15.3 ppm. The comparison between the CPMAS and MAS spectra shows a different intensity for this methyl signal reflecting the greater mobility of this group, as expected. Paradoxically, the intensity of the carbonyl signal does not increase in the MAS spectrum compared with the
signal of the methyl group of CD. Finally, no significant differences were obtained between the non-activated polyBTCA-CD and the activated polyBTCA-CD spectra.

Metal adsorption
Effect of activation (NaHCO$_3$ treatment)
The activated polymer is more efficient than the non-activated polymer regardless of the dose (Figure 5). The activation enhanced the removal by 69%, 92%, 78%, 92% and 92% at a polymer concentration of 2 g·L$^{-1}$ for Al$^{3+}$, Co$^{2+}$, Cr$^{3+}$, Ni$^{2+}$ and Zn$^{2+}$, respectively. Moreover, it appears that a polymer concentration of 2 g·L$^{-1}$ is enough to treat an inorganic load of 50 mg·L$^{-1}$. Thus, for all the following experiments, the polymers were activated and a concentration of 2 g·L$^{-1}$ was used.

Adsorption kinetics
Figure 6 shows the adsorption kinetics for two solutions containing five metals at 1 mg·L$^{-1}$ and 10 mg·L$^{-1}$. For these two concentrations, 100% removal was reached for most species except for Al$^{3+}$ and Cr$^{3+}$ systems, which reached a state of dynamic equilibrium. However, the adsorption time changed with the concentration and with the metals. At 1 mg·L$^{-1}$ the adsorption kinetics was fast: in 5 min equilibrium was reached for the metals except for Al$^{3+}$ (240 min). At 10 mg·L$^{-1}$ for
Co²⁺, Ni²⁺ and Zn²⁺, equilibrium time increased to 30 min, while for Al³⁺ and Cr³⁺ it decreased to 30 min. These results were obtained in triplicate with small standard deviations, indicating the reproducibility of the experiments. For the two experiments we noted an increase of pH values ranging from 3.9 and 4.4 for the initial pH (pHᵢ) to 6.2 and 7.3 for the final pH (pHᶠ) for concentrations of 10 mg·L⁻¹ and 1 mg·L⁻¹, respectively.

Influence of metal and calcium concentration
The results show that activated polyBTCA-CD is able to treat metal solutions from low concentrations (a few micrograms in SS3 and SS4 in Table 1) to high concentrations (ten milligrams for each metal (SS1 in Table 1). The exact composition of each spiked solution (SS) is given in the Experimental section. In addition, it was observed that without Ca²⁺ cations the removal efficiency was above 99%, except for aluminum and chromium (Table 1). When expressing these results in mmol of total metal retained per gram of polymer, it can be noted that for the highest concentration of metals (10 mg·L⁻¹), the polymer retained 0.466 mmol·g⁻¹. When calcium ions were added (SS5 in Table 1), Cr³⁺ was better retained than in the solution without calcium (85% compared to 38%) while the retention of Al³⁺ remained constant. For the three other metals, the retention dramatically decreased (<41%) in the presence of calcium.

PAH and AP adsorption
Figure 7a shows that the polyBTCA-CD can take up PAHs since the polymer removed between 74% and 79% of the global PAH load. It was also observed that removal was more efficient for heavy PAHs (89% adsorption) than for the lighter ones (62% adsorption). Following PAH adsorption the pH ranged from 6 to 7.6 at the end of the experiment. The polymer successfully removed APs as shown in Figure 7b. It can be noted that an increase of AP concentration led to a decrease in adsorp-
Table 1: Efficiency of activated polyBTCA-CD expressed in % (concentration = 2 g·L⁻¹) to treat several metal spiked solutions SS (n = 3). See Experimental section for the exact compositions of the spiked solutions SS.

<table>
<thead>
<tr>
<th></th>
<th>SS1</th>
<th>SS2</th>
<th>SS3</th>
<th>SS4</th>
<th>SS5</th>
</tr>
</thead>
<tbody>
<tr>
<td>Al³⁺</td>
<td>87 ± 4</td>
<td>66 ± 4</td>
<td>87 ± 16</td>
<td>95 ± 2</td>
<td>97 ± 0</td>
</tr>
<tr>
<td>Co²⁺</td>
<td>&gt;99</td>
<td>&gt;99</td>
<td>&gt;99</td>
<td>&gt;99</td>
<td>—</td>
</tr>
<tr>
<td>Cr³⁺</td>
<td>88 ± 3</td>
<td>36 ± 1</td>
<td>45 ± 7</td>
<td>38 ± 2</td>
<td>17 ± 2</td>
</tr>
<tr>
<td>Ni²⁺</td>
<td>&gt;99</td>
<td>&gt;99</td>
<td>99 ± 0</td>
<td>99 ± 0</td>
<td>—</td>
</tr>
<tr>
<td>Zn²⁺</td>
<td>&gt;99</td>
<td>&gt;99</td>
<td>&gt;99</td>
<td>—</td>
<td>41 ± 7</td>
</tr>
<tr>
<td>Ca²⁺</td>
<td>—</td>
<td>—</td>
<td>—</td>
<td>—</td>
<td>10 ± 4</td>
</tr>
<tr>
<td>pHf</td>
<td>6.2</td>
<td>7.3</td>
<td>7.3</td>
<td>7.3</td>
<td>4.7</td>
</tr>
</tbody>
</table>

Table 2: Removal efficiency expressed in % of metals, PAHs and APs in solutions containing either one family of substances or mixtures after treatment by polyBTCA-CD (concentration = 2 g·L⁻¹, n = 3).

<table>
<thead>
<tr>
<th></th>
<th>metals</th>
<th>PAHs</th>
<th>APs</th>
<th>metals + PAHs</th>
<th>metals + APs</th>
<th>PAHs + APs</th>
<th>metals + PAHs + APs</th>
</tr>
</thead>
<tbody>
<tr>
<td>Al³⁺</td>
<td>66 ± 4</td>
<td>—</td>
<td>—</td>
<td>65 ± 14</td>
<td>63 ± 5</td>
<td>—</td>
<td>65 ± 5</td>
</tr>
<tr>
<td>Co²⁺</td>
<td>&gt;99</td>
<td>—</td>
<td>—</td>
<td>99 ± 1</td>
<td>99 ± 1</td>
<td>—</td>
<td>99 ± 1</td>
</tr>
<tr>
<td>Cr³⁺</td>
<td>36 ± 1</td>
<td>—</td>
<td>—</td>
<td>28 ± 8</td>
<td>25 ± 13</td>
<td>—</td>
<td>24 ± 13</td>
</tr>
<tr>
<td>Ni²⁺</td>
<td>&gt;99</td>
<td>—</td>
<td>—</td>
<td>98 ± 2</td>
<td>98 ± 1</td>
<td>—</td>
<td>98 ± 1</td>
</tr>
<tr>
<td>Zn²⁺</td>
<td>&gt;99</td>
<td>—</td>
<td>—</td>
<td>96 ± 4</td>
<td>96 ± 4</td>
<td>—</td>
<td>96 ± 4</td>
</tr>
<tr>
<td>PAHs</td>
<td>light</td>
<td>60 ± 7</td>
<td>—</td>
<td>63 ± 3</td>
<td>—</td>
<td>62 ± 8</td>
<td>62 ± 4</td>
</tr>
<tr>
<td></td>
<td>heavy</td>
<td>87 ± 8</td>
<td>—</td>
<td>81 ± 10</td>
<td>—</td>
<td>87 ± 3</td>
<td>89 ± 5</td>
</tr>
<tr>
<td>APs</td>
<td>4NP</td>
<td>—</td>
<td>—</td>
<td>80 ± 3</td>
<td>82 ± 7</td>
<td>80 ± 11</td>
<td>81 ± 1</td>
</tr>
<tr>
<td></td>
<td>4nNP</td>
<td>—</td>
<td>—</td>
<td>89 ± 1</td>
<td>90 ± 2</td>
<td>85 ± 1</td>
<td>81 ± 2</td>
</tr>
<tr>
<td></td>
<td>4tOP</td>
<td>—</td>
<td>—</td>
<td>83 ± 3</td>
<td>86 ± 6</td>
<td>86 ± 6</td>
<td>82 ± 4</td>
</tr>
<tr>
<td>pHf</td>
<td>7.3</td>
<td>7.6</td>
<td>8</td>
<td>7.7</td>
<td>7.7</td>
<td>7.7</td>
<td>7.9</td>
</tr>
</tbody>
</table>
Table 3: Average concentrations expressed in mg·L\(^{-1}\) and standard errors of the main elements present in the DWs (\(n = 5\)).

<table>
<thead>
<tr>
<th>Element</th>
<th>Concentration (mg·L(^{-1})) ± Standard Error</th>
</tr>
</thead>
<tbody>
<tr>
<td>Al(^{3+})</td>
<td>1.48 ± 0.54</td>
</tr>
<tr>
<td>Co(^{2+})</td>
<td>1.70 ± 0.74</td>
</tr>
<tr>
<td>Cr(^{3+})</td>
<td>0.04 ± 0.03</td>
</tr>
<tr>
<td>Ni(^{2+})</td>
<td>0.25 ± 0.11</td>
</tr>
<tr>
<td>Zn(^{2+})</td>
<td>0.90 ± 0.50</td>
</tr>
<tr>
<td>Ca(^{2+})</td>
<td>690 ± 156</td>
</tr>
<tr>
<td>Fe(^{3+})</td>
<td>0.23 ± 0.12</td>
</tr>
<tr>
<td>K(^+)</td>
<td>73.3 ± 6.56</td>
</tr>
<tr>
<td>Mg(^{2+})</td>
<td>2.84 ± 0.31</td>
</tr>
<tr>
<td>Mn(^{2+})</td>
<td>0.12 ± 0.1</td>
</tr>
<tr>
<td>Sr(^{2+})</td>
<td>0.24 ± 0.03</td>
</tr>
</tbody>
</table>

Detailed analysis of the discharge water after treatment with polyBTCA-CD

Among the 189 substances/parameters analyzed in the raw DW, 35 were detected: sixteen inorganic elements including twelve metals, and seven APs (Table 5). The treatment by activated polyBTCA-CD affected all water parameters and all substances detected in the raw DW except for Na. As in previous experiments on DWs, it was observed that the metals were not efficiently adsorbed by the polymer with a removal efficiency reaching 24%, 12%, 17% and 44% for Al\(^{3+}\), Co\(^{2+}\), Ni\(^{2+}\) and Zn\(^{2+}\), respectively. Unlike metals, it appears that most of the organics were retained with removal efficiencies higher than 50%; for example 4tOP, 4nNP, monoethoxylate nonylphenol and monoethoxylate octylphenol were efficiently removed (81, 66, 83 and 77%, respectively). Moreover, some substances were retained in large quantities on the polyBTCA-CD such as chloride (130 mg·L\(^{-1}\)) and calcium (119 mg·L\(^{-1}\)), which is not reflected by the removal efficiency because of the large amounts present in the initial DW. The sodium concentration remained stable after adsorption due to the fact that the material was initially activated by NaHCO\(_3\).

Discussion

To explain the adsorption performance of polyBTCA-CD, a chemisorption mechanism involving several interactions can occur including ion exchange, electrostatic interactions, inclusion complexation and/or precipitation [3,12-16].
The polymer, without an activation step, removed 20% of the total metal load in spiked solutions (SS) containing five metals at 10 mg·L$^{-1}$ each. This could be attributed to surface adsorption and diffusion into the polymer network. However, it was observed that an activation step in an aqueous solution of sodium bicarbonate was necessary to enhance metal retention, as also reported by Ducoroy et al. [17], allowing for a removal efficiency for metals higher than 99% for Co$^{2+}$, Ni$^{2+}$ and Zn$^{2+}$, through both electrostatic interactions and ion exchange. Moreover, this treatment only converted carboxylic groups into carboxylate without altering the amorphous structure of the polymer, as seen from the X-ray spectra.

Precipitation could also explain the removal efficiency enhancement observed with the activated polymer since it induced a strong pH increase. Indeed, for the solution of five metals at concentration of 1 mg·L$^{-1}$ (SS2) treated with activated

<table>
<thead>
<tr>
<th>Table 5: Extensive analysis on raw and treated DW by polyBTCA-CD (concentration = 2 g·L$^{-1}$).</th>
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</thead>
<tbody>
<tr>
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<tr>
<td>------------------</td>
</tr>
<tr>
<td>physicochemical parameters (mg L$^{-1}$)</td>
</tr>
<tr>
<td>pH</td>
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<tr>
<td>BOD-5</td>
</tr>
<tr>
<td>COD</td>
</tr>
<tr>
<td>hydrocarbon index C10-C40</td>
</tr>
<tr>
<td>total cyanide</td>
</tr>
<tr>
<td>AOX</td>
</tr>
<tr>
<td>nitrates</td>
</tr>
<tr>
<td>total nitrogen</td>
</tr>
<tr>
<td>inorganic elements (mg·L$^{-1}$)</td>
</tr>
<tr>
<td>chloride</td>
</tr>
<tr>
<td>sulphate</td>
</tr>
<tr>
<td>potassium</td>
</tr>
<tr>
<td>calcium</td>
</tr>
<tr>
<td>magnesium</td>
</tr>
<tr>
<td>manganese</td>
</tr>
<tr>
<td>sodium</td>
</tr>
<tr>
<td>sulfur</td>
</tr>
<tr>
<td>aluminum</td>
</tr>
<tr>
<td>cobalt</td>
</tr>
<tr>
<td>iron</td>
</tr>
<tr>
<td>molybdenum</td>
</tr>
<tr>
<td>nickel</td>
</tr>
<tr>
<td>selenium</td>
</tr>
<tr>
<td>strontium</td>
</tr>
<tr>
<td>zinc</td>
</tr>
<tr>
<td>organic substances (µg·L$^{-1}$)</td>
</tr>
<tr>
<td>1,2-dichloroethane</td>
</tr>
<tr>
<td>chloroform</td>
</tr>
<tr>
<td>dichlorobromomethane</td>
</tr>
<tr>
<td>4-tert-octylphenol</td>
</tr>
<tr>
<td>4-nonylphenol</td>
</tr>
<tr>
<td>monoethoxylate nonylphenol</td>
</tr>
<tr>
<td>diethoxylate nonylphenol</td>
</tr>
<tr>
<td>monoethoxylate octylphenol</td>
</tr>
<tr>
<td>diethoxylate octylphenol</td>
</tr>
<tr>
<td>4-tert-butylphenol</td>
</tr>
</tbody>
</table>
polyBTCA-CD, the pH increased from 4.4 to 7.3, probably due to the basic character of the polymer, the COOH groups of which have been converted into COO$^-$/Na$^+$ groups inducing metal precipitation, notably for Al$^{3+}$ and Cr$^{3+}$.

With the highest concentration tested (10 mg L$^{-1}$, which corresponds to 5 mg of each metal per gram of polymer), the removal efficiency represents 0.466 mmol of metals adsorbed per gram of polymer, that is lower than the theoretical ion exchange capacity, estimated by titration to be 0.705 mmol of COOH functional groups per gram of polymer. In theory, in order to saturate the COOH functional groups, a concentration of 6.5 mg of each metal per gram of polymer (i.e., 13 mg L$^{-1}$) would have been necessary. In the DWs, focusing on Ca$^{2+}$ only, 119 mg L$^{-1}$ were retained by the polymer, representing a concentration of 1.49 mmol g$^{-1}$ of material. Thus, the IEC of the polymer (0.705 mmol g$^{-1}$) was exceeded. This observation confirmed the fact that the inorganic elements were not only retained by ion exchange and/or electrostatic interactions, but other interactions occurred, such as precipitation and/or physical phenomena (surface adsorption, diffusion into the network and/or hydrogen bonding).

This study demonstrated the efficiency of activated polyBTCA-CD to treat the inorganic load, but also to treat the organic pollutants. It can be noted that the analysis of 26 substances present in the DW tested, showed that the polymer is able to adsorb 4-tert-butylphenol, 1,2-dichloroethane and chloroform, but also APs and the organo-halogenated compounds (represented by the AOX parameter). For the retention of APs, no difference was noted between 4NP, 4nNP and 4OP. Heavy PAHs were better retained than light ones. This observation could be partially explained by the greater hydrophobicity of heavy PAHs compared with light ones [19]. Indeed, in the case of interactions with organics, CD units play an important role. The hydrophobic cavities of these molecules allow for the formation of inclusion complexes with PAH guest molecules (organic substances) [3,20]. If the guest compounds present an appropriate molecular size and structure to enter into the CD cavity the more hydrophobic the organic substances, the more stable the inclusion complexes.

The material was also able to efficiently treat complex solutions containing metals, PAHs, APs and other substances. However, although in spiked solutions no competition effect was revealed between substance groups, these phenomena appeared in more complex solution such as DWs. Indeed, the retention of inorganic elements decreased from an average of 87% for the five metals in SS to 26% in DWs. It appears that other elements present in DW were retained by the polymer, for instance Fe$^{3+}$, Mn$^{2+}$, Mg$^{2+}$ and Ca$^{2+}$. Discharge waters from ST industries are complex matrices containing not only metals, PAHs or APs but also other elements including anions and salts, sometimes exceeding one gram per liter [21]. In this study, calcium and chlorides were present in the DWs at concentrations of 500 and 3 000 mg L$^{-1}$, respectively. In fact, a large amount of these other inorganic elements could interact with the polymer, notably Ca$^{2+}$, which was retained at 119 mg L$^{-1}$. Thus, the low retention of target metals (Al$^{3+}$, Co$^{2+}$, Cr$^{3+}$, Ni$^{2+}$ and Zn$^{2+}$) by the adsorbent could be due to the simultaneous presence of high quantities of K$^+$, Na$^+$, Mg$^{2+}$ or Ca$^{2+}$, which can saturate the carboxylate functions of the polymer and compete with the target metals for access to the active sites [22-25]. These observations were confirmed by the tests conducted in SS in the presence of calcium. An increase of polymer concentration in DWs (10 g L$^{-1}$ instead of 2 g L$^{-1}$) yielded better results in terms of removal efficiency for all metals. However, while the retention of Cr$^{3+}$, Ni$^{2+}$ and Zn$^{2+}$ was reduced in the presence of Ca$^{2+}$, this was not always the case for Al$^{3+}$ and Cr$^{3+}$. In some cases, Al$^{3+}$ and Cr$^{3+}$ exhibited unexpected behavior.

In solutions at low concentrations (SS2, SS3 and SS4), Al$^{3+}$ and Cr$^{3+}$ were removed with low efficiency compared to other metals, but in solutions at higher concentrations (SS1) or in the presence of calcium (SS5), an increase of removal efficiency was observed whereas for the three other metals, retention decreased. In the case of mixtures, the same observations were made, Al$^{3+}$ and Cr$^{3+}$ removal being systematically lower, whereas for the three other metals the removal efficiency remained as high as 99%. Moreover, in DWs, Al$^{3+}$ was retained by an average of 60% while the others did not exceed a removal efficiency of 30%. No simple explanation could clarify these phenomena but some specific characteristics of these two metals must be pointed out. Indeed, the behavior of both Al$^{3+}$ and Cr$^{3+}$ is highly pH dependent, with complex speciation patterns [26-28]. Since the pH value was not kept constant during tests, the pH variations probably led to changes in speciation of the two metals, leading to this unexpected behavior compared to the other metals. Moreover, it can be noted that Al$^{3+}$ and Cr$^{3+}$ have the smallest ionic radii of the five metals and a low electronegativity, which are two parameters that decrease metal removal as described by Al-Rub and co-workers [29].

**Conclusion**

Throughout this study, it was demonstrated that polyBTCA-CD is a versatile sorbent able to retain substances present at concentrations close to a few milligrams per liter (metals and other inorganic elements) but also at trace concentrations (μg L$^{-1}$ for organics). Although ion exchange on the one hand, and host–guest inclusion on the other hand were the main phenomena interacting between adsorbent and solutes, the interpretation...
of the results was made difficult due to the wide diversity of polluting species present in DWs, involving numerous other mechanisms in the adsorption (classified for instance as chemisorption and physisorption phenomena) [3,12,20]. Thus, further studies are needed to better understand how water parameters impact the different routes of pollutant retention, which is of great interest for future applications of non-conventional adsorbents in industry.

Experimental
Synthesis and characterization of the material
The cross-linked polymer (Figure 1) was prepared in a single step by cross-linking hydroxypyropyl-β-cyclodextrin (HPβCD; Kleptose HPB®, DS = 0.62, Roquette, Lestrem, France) using 1,2,3,4-butanetetraacrylic acid (BTCA; Aldrich, Milwaukee, WI). The synthesis procedure has already been described in detail by Martel et al. [10,30]. Based from their methodology, the typical cross-linking reaction was carried out as follows: in a reactor, 0.37 mol·L⁻¹ of sodium dihydrogenophosphate (Na₂H₂PO₄·H₂O, catalyst), 0.85 mol·L⁻¹ of BTCA and 0.18 mol·L⁻¹ of HPβCD were dissolved; the resulting solution was then concentrated by evaporation, and the mixture heated to 140 °C. These last two steps were performed under vacuum in a rotary evaporator; the polymer formed was re-suspended through addition of water to the reactor; the polymer was then filtered and purified by several washings with water. It was dried at 50 °C until constant weight, yielding a white powder. The yield of the reaction was equal to 87.6 %.

To activate the polyBTCA-CD by converting the carboxylic functions to their carboxylate form, the material was stirred for 4 h in an aqueous solution of 4 g·L⁻¹ NaHCO₃ (saturated) then extensively washed with osmosed water to remove unreacted reagents, and finally oven dried at 50 °C.

The ion exchange capacity (IEC) of the material was determined by pH-metric titration according to the calcium acetate method [31]. A solution of calcium acetate (2% w/w) was prepared in osmosed water. A weighed sample of dry polymer was stirred into the calcium acetate solution for 2 h following calibration with oxalic acid. Then, the solution was filtered, and the amount of acetic acid formed was measured by titration using a sodium hydroxide solution (0.05 M). The amount of ion exchange functions was equal to the amount of acetic acid present in solution. The results were expressed in mmol of COOH functional groups per gram of polymer.

The point of zero charge (PZC) values of both non-activated and activated polyBTCA-CD were determined by titration with the salt addition method. 50 mL of 0.1 M NaCl solution was placed in each of nine beakers. The pH of each solution was adjusted from 3 to 11 with one pH unit increment using a pH meter (pH meter, model 3110, WTW, Alès, France) with aqueous 0.1 M NaOH and 0.1 M HCl. Then 0.15 g of polymer was added to each beaker. The solutions were stirred for 48 h to reach equilibrium then the resulting pH (pHf) was measured. The difference between the initial and final pH values (ΔpH) was plotted against the initial pH. The PZC was represented by the point ΔpH = 0 [32].

X-ray diffraction analysis (XRD) was performed on a Bruker D8 Advance diffractometer using Cu Kα radiation with a wavelength of λ = 0.15406 nm produced at 40 kV and 40 mA. XRD data were collected over the 20 range from 10° to 60° at every 0.02° with a scan speed of 0.5 s per step. For polymer, the diffraction profiles are divided into 2 parts: peaks related to diffraction of crystallites and a broad peak related to scattering of the amorphous phase. The assumption is that the areas are proportional to the scattering intensities of crystalline and amorphous phases. Thus, the percentage of polymer that is crystalline was determined from Equation 1 [33]:

\[
\% \text{ crystallinity } = \frac{\text{area under crystalline peaks}}{\text{total area under all peaks}} \times 100. \quad (1)
\]

The polymer was also characterized by solid-state ¹³C NMR techniques such as cross polarization magic angle spinning (CP MAS) and MAS. The spectra were recorded with a Bruker spectrometer operating at 75.47 MHz and 303 K. The compounds were placed in a zirconium rotor, 4 mm in diameter and 21 mm high. The chemical shifts were recorded relative to tetramethylsilane with benzene as secondary reference. The Hartmann–Hahn condition was satisfied during cross polarization magic angle with 1.5 ms of contact time under the following conditions: repetition time 8 s, 1H90° pulse length 4 µs, and spin rate at 10 kHz.

Adsorption tests
To determine the ability of the material to treat inorganic and organic load, several batch experiments were carried out with two kinds of solutions: spiked solutions (SS) containing different substances at several concentrations that are typical of ST discharge waters (DWs) [34] and real ST industrial DWs themselves.

SS contained metal sulfate salts (Al₂(SO₄)₃·16H₂O; CoSO₄·7H₂O; Cr₂(SO₄)₃·12H₂O; NiSO₄·6H₂O; ZnSO₄·7H₂O) purchased from Aldrich and used without further purification. The sixteen PAHs of the US EPA list [35] and three APs were purchased from Supelco Sigma Aldrich (Saint-Quentin Fallavier, France) and used as received; eight light PAHs: naph-
thelene (NAP), acenaphthene (ACE), acenaphthylene (ACY), fluorene (FLU), phenanthrene (PHE), anthracene (ANT), fluoranthene (FLT) and pyrene (PYR); eight heavy PAHs: benzo[a]anthracene (BaANT), chrysene (CHY), benzo[ghi]fluoranthene (BghiFLT), benzo[a]pyrene (BaPYR), dibenz[a,h]anthracene (dBahANT), indeno[1,2,3-cd]pyrene (IcdPYR) and benzo[g,h,i]perylene (BghiPL); three APs: 4-nonylphenol (4NP, CAS no. 84852-15-3), 4-n-nonylphenol (4nNP, CAS no. 104-40-5) and 4-tert-octylphenol (4tOP, CAS no. 140-66-9). Calcium chloride and sodium bicarbonate were purchased from Fischer Scientific (Illkirch, France) and used as received. Each SS was prepared from stock solutions in osmosed water.

### Metal adsorption
In order to examine the effect of NaHCO₃ treatment (activation), the adsorption capacities of the non-activated and the activated polymer polyBTCA-CD were compared with a solution containing the five metals (Al³⁺, Co²⁺, Cr³⁺, Ni²⁺ and Zn²⁺) each at a concentration of 10 mg·L⁻¹. A kinetic study was also performed to determine the appropriate contact time between the material and the polymeric solutions containing concentrations of 1 and 10 mg·L⁻¹ for each metal. Then, metal adsorption was determined by several polymetallic solutions containing the five metals described previously. Different concentrations of metals were tested: 10 mg·L⁻¹ (SS1) and 1 mg L⁻¹ (SS2) for each metal and two solutions reproducing DW concentrations (SS3 and SS4). Moreover, two experiments were performed with metals in the absence (SS4) or presence (SS5) of CaCl₂ at DW concentrations, in order to observe the influence of salt concentrations (Table 6).

### Adsorption capacities in discharge waters
Five samples of real discharge waters (DWs) were collected from Zindel Industry located in Seloncourt (Doubs, France) which is specialized in chemical coatings and any processes for the corrosion protection of metal parts intended for the automotive and building sectors. Their process waters mainly contain metallic pollutants (e.g., Cr³⁺ and Zn²⁺) coming from rinsing and washing baths. Following the tests conducted on SS, metal retention capacity was also tested on five different industrial DWs. In one DW, the effect of polymer dose was tested: 5, 10, 15 and 20 g·L⁻¹. Moreover, one more detailed analysis was performed in order to see if the polymer can retain other substances present in the DW.

### Batch experiments
In each experiment, 2 g·L⁻¹ of material was stirred (250 rpm) with a fixed volume of polluted solution (with no modification of the initial pH value) at room temperature for 4 h. After treatment, the solutions were left to settle for 1 h, and the supernatant was analyzed. Chemical analyses were performed in the initial and in the treated solutions and the results expressed in concentration and removal efficiency. Under the same conditions, a control experiment was performed without pollutants in order to check whether pH variations occurred in non-activated and activated polyBTCA-CD.

### Chemical analyses
For each SS or DW, initial and final pH values were measured. Metal concentrations were determined by inductively coupled plasma atomic emission spectroscopy (ThermoFisher, iCAP
References


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The definitive version of this article is the electronic one which can be found at: doi:10.3762/bjoc.12.172