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## Extraction of $\alpha$ -mangostin from *Garcinia mangostana* L. using alternative solvents: Computational predictive and experimental studies

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### A B S T R A C T

This study evaluated the performance of alternative green solvents, i.e. *D*-limonene, dimethylcarbonate (DMC), ethanol, ethyl acetate, ethyl lactate and methyltetrahydrofuran (MeTHF) compared to the petroleum based dichloromethane, for extraction of  $\alpha$ -mangostin from *Garcinia mangostana* pericarps. The Hansen solubility parameters (HSPs) were used to explain the dissolution behavior of the solutes and solvents, and the conductor-like screening model for realistic solvation. The (COSMO-RS), a statistical thermodynamic approach based on the results of quantum chemical calculations for comprehending the dissolving mechanisms were used to predict the extraction prediction. On the basis of the Hansen analysis, dichloromethane was the most suitable solvent for extraction of  $\alpha$ -mangostin. However, COSMO-RS analysis showed a higher solubility of  $\alpha$ -mangostin in ethyl lactate, DMC, MeTHF, ethyl acetate and ethanol. Moreover, the experimental studies using a classical reflux extraction followed by a quantitative HPLC analysis of  $\alpha$ -mangostin showed similar results to the predictive values from the COSMO-RS model. The  $\alpha$ -mangostin levels extracted by ethyl lactate, DMC, MeTHF, ethyl acetate and ethanol were higher than those using dichloromethane and *D*-limonene. The results support the potential use of ethyl lactate, DMC, MeTHF, and ethanol as alternative green solvents for the preparation of  $\alpha$ -mangostin extracts.

#### Keywords:

Green solvents  
Extraction  
*Garcinia mangostana*  
Hansen  
COSMO-RS

### 1. Introduction

*Garcinia mangostana* L. (mangosteen) is an important economic plant grown in Eastern and Southern Thailand. The fruits are popularly known as the “queen of fruits”, as they produce a popular refreshing tasty juice. The fruit pericarps, are also known for their ethno-medicinal values and these are a by-product of the mangosteen juice factory. *G. mangostana* pericarp extracts are enriched in xanthenes, a group of bioactive compounds (Zarena &

Sankar, 2011). The major compounds in the extracts are prenylated xanthenes, particularly  $\alpha$ -mangostin (Nualkaew et al., 2012). This has been used as a standard marker for their standardization. Extracts from *G. mangostana* pericarps have been reported for their various pharmacological activities, including their actions as antioxidants, and their anticancer (Moongkarndi et al., 2014), antimicrobial (Sakagami, Iinuma, Piyasena, & Dharmaratne, 2005), anti-inflammatory (Cui et al., 2010), and wound healing (Nganlasom, Suttitum, Jirakulsomchok, & Puapairoj, 2008) properties. The extracts of *G. mangostana* pericarps have also been used in several commercial nutritional supplements, pharmaceutical products and cosmetics (Pothitirat, Chomnawang, & Gritsanapan, 2009). It has been reported that dichloromethane was the most suitable solvent for extraction of  $\alpha$ -mangostin from *G. mangostana* pericarps (Pothitirat, Chomnawang, & Gritsanapan, 2010). However, the use

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of dichloromethane, a halogenated hydrocarbon as a solvent for extraction is restricted due to its toxicities (CalEPA, 2000), and limitations on its use in the development of herbal products.

Nowadays, industrial regulations require a diminished consumption of petrochemical solvents and volatile organic compounds. Moreover, the absence of risk during extraction and the safety of the ingredients used are a major concern and have drawn attention towards the need to use greener solvents (Rombaut, Fabiano-Tixier, Bily, & Chemat, 2014). Alternative green solvents, a renewable resource produced from biomasses such as wood, starch, vegetable oils or fruits are high power, biodegradable, low toxicity and low flammability. Extensive studies have shown the importance and the reported potential of green solvents that could be alternatively used instead of the petrochemical solvents, e.g. extraction of fat and oils from rapeseed using MeTHF instead of hexane (Sicaire et al., 2015), and extraction of oil from microalgae using *D*-limonene instead of hexane (Tanzi, Abert-Vian, Ginies, Elmaataoui, & Chemat, 2012). Recently, the dissolving power and selectivity of solvents can be theoretically predicted using computational programs. Currently two computational programs, i.e. the Hansen solubility parameters (HSPs) and a conductor like screening model for realistic solvation (COSMO-RS) are being used to predict the most suitable solvents for extraction of natural products (Filly et al., 2014; Filly, Fabiano-Tixier, Fernandez, & Chemat, 2015; Li, Fabiano-Tixier, Ginies, & Chemat, 2014).

The aim of this work was to evaluate the ability of some alternative 'green' solvents to extract  $\alpha$ -mangostin compared to dichloromethane using the computational programs and experimental extractions (Fig. 1). The computational predictive methods were performed via an HSP program that explained the dissolution behavior of solutes and solvents, and the COSMO-RS program, a statistical thermodynamics approach based on the result of quantum chemical calculations that provided explanations of the mechanisms of dissolution. The results from these predictive computational methods were confirmed by those from experimental extractions.

## 2. Materials and methods

### 2.1. Plants material

Fresh mature *G. mangostana* fruits were from a market in the Hat-Yai district, Songkhla province, Thailand. The fruits of *G. mangostana* were cleaned and the edible parts were removed. The pericarps were cut into small pieces, and dried at 50 °C for 72 h in a hot air oven. The dried pericarps were then ground into a fine powder using a grinder, and passed through a sieve No. 45. The dried pericarp powder was kept in a well-closed container and

protected from light until further use.

### 2.2. Reagents

The solvents used were analytical grade. Dichloromethane (purity 99.9%) was from Merck, Germany. Ethyl acetate (purity 99%) and (R)-(+)-limonene (purity 93%) were from Sigma Aldrich, Germany. Ethanol (purity 100%) and 2-methyltetrahydrofuran (2-MeTHF) (purity 99%) were from VWR International, Germany. Ethyl lactate (purity 99%) and dimethylcarbonate (DMC) (purity 99%) were from Alfa Aesar, Germany. A sample of  $\alpha$ -mangostin was purified and identified by the method described by Pothitirat et al. (2010).

### 2.3. Computational methods: HSPs

Solubility parameters for predicting the solubility of a solute was proposed by Hansen (Hansen, 1969). HSPs are based on the concept that the total cohesive energy density is approximated by the sum of the energy densities required to overcome atomic dispersion forces ( $\delta_d^2$ ), molecular polar forces arising from dipole moments ( $\delta_p^2$ ) and hydrogen-bonds between molecules ( $\delta_h^2$ ), as given in the following equation.

$$\delta_{\text{total}}^2 = \delta_d^2 + \delta_p^2 + \delta_h^2 \quad (1)$$

Where  $\delta_{\text{total}}$  is the Hansen total solubility parameter, which consists of three HSPs in terms of dispersion ( $\delta_d$ ), polar ( $\delta_p$ ) and hydrogen bonding ( $\delta_h$ ).

For HSP solvent optimization, a simple composite affinity parameter, the relative energy difference (RED) number, has been calculated using Equation (2) to determine the solubility between solvent and solute.

$$\text{RED} = R_a/R_o \quad (2)$$

Where  $R_o$  is the radius of a Hansen solubility sphere, and  $R_a$  is the distance of a solvent from the center of the Hansen solubility sphere, given by Equation (3):

$$R_a^2 = 4(\delta_d A - \delta_d B)^2 + (\delta_p A + \delta_p B)^2 (\delta_h A - \delta_h B)^2 \quad (3)$$

Where A refers to the solute and B refers to the solvent.

The factor 4 in Eq. (3), based on Prigogine's Corresponding States Theory, has proved to effectively expand the dimensions in order to give spherical plots. In general, the parameter follows the classical "like dissolves like" rule: the smaller  $R_a$  is the greater affinity between solute and solvent. It means that good potential solvents have an RED of less than 1 (the compound has similar properties

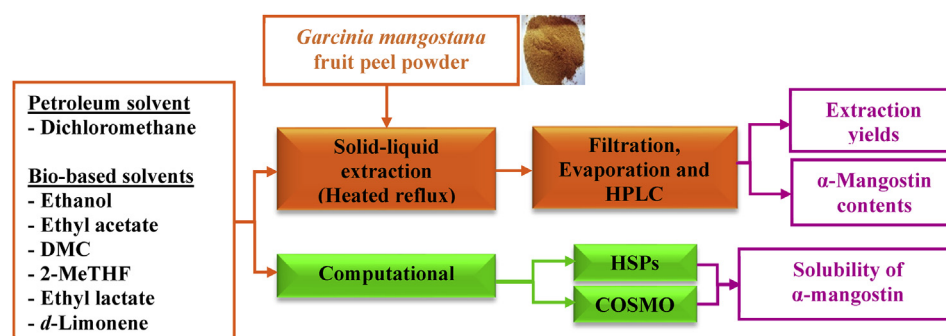


Fig. 1. Experimental design and diagram for extraction of *G. mangostana* fruit peels. DMC, dimethylcarbonate, 2-MeTHF, 2-methyltetrahydrofuran. HPLC, high performance liquid chromatography. HSPs, Hansen solubility parameters. COSMO, conductor like screening model for realistic solvation

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and will dissolve), while medium and poor solvents have RED values of from 1 to 3 and more than 3, respectively. The chemical structures of the solvents and solute (Fig. 2) could be mutually transformed by the JChemPaint version 3.0.1 software, then to their simplified molecular input line entry syntax (SMILES) notations, which were subsequently used for the HSP calculation using the Yamamoto–Molecular Break (Y-MB) method, which can break HSPs into corresponding functional groups and then estimate its HSPs. This method has been embedded in the HSPiP software in order to facilitate the direct HSP calculation for other compounds. These solubility parameters were further modeled to a three-dimensional HSP sphere for better visualizing the solute/solvent interactions (HSPiP Version 4.0, Denmark). This method has been embedded in the HSPiP software in order to facilitate the calculation of the HSP of  $\alpha$ -mangostin and solvents.

#### 2.4. Computational methods: COSMO-RS

COSMO-RS uses a statistical thermodynamics approach based on the result of quantum chemical calculations for an understanding of the dissolving mechanism. COSMO-RS combines quantum chemical considerations (COSMO) and statistical

thermodynamics (RS) to determine and predict the thermodynamic properties. The COSMO-RS developed by Klamt (2003) is known as a powerful method for molecular description and solvent screening based on a quantum-chemical approach.

COSMO-RS is a two-step procedure including microscopic and macroscopic steps. In the first step, the COSMO model is applied to simulate a virtual conductor environment for the molecule. The molecule is embedded into a virtual conductor (Fig. 3a). In such an environment, the molecule induced a polarization charge density on its surface (Fig. 3b). Therefore, during the quantum calculation self-consistency algorithm the solute molecule is converged to its energetically optimal state in the conductor with respect to its electron density and geometry. The standard quantum chemical methods, triple zeta valence polarized basis set (TZVP) was used in this study. The second step used the statistical thermodynamic calculation. This polarization charge density was used for the quantification of the interaction energy of the pair-wise interacting surface segments with regard to electrostatics and hydrogen bonding. The 3D distribution of the polarization charges on the surface of each molecule was converted into a surface composition function ( $\sigma$ -profile) that provided information of the distribution of the molecular polarity (Fig. 3c). The thermodynamics of the

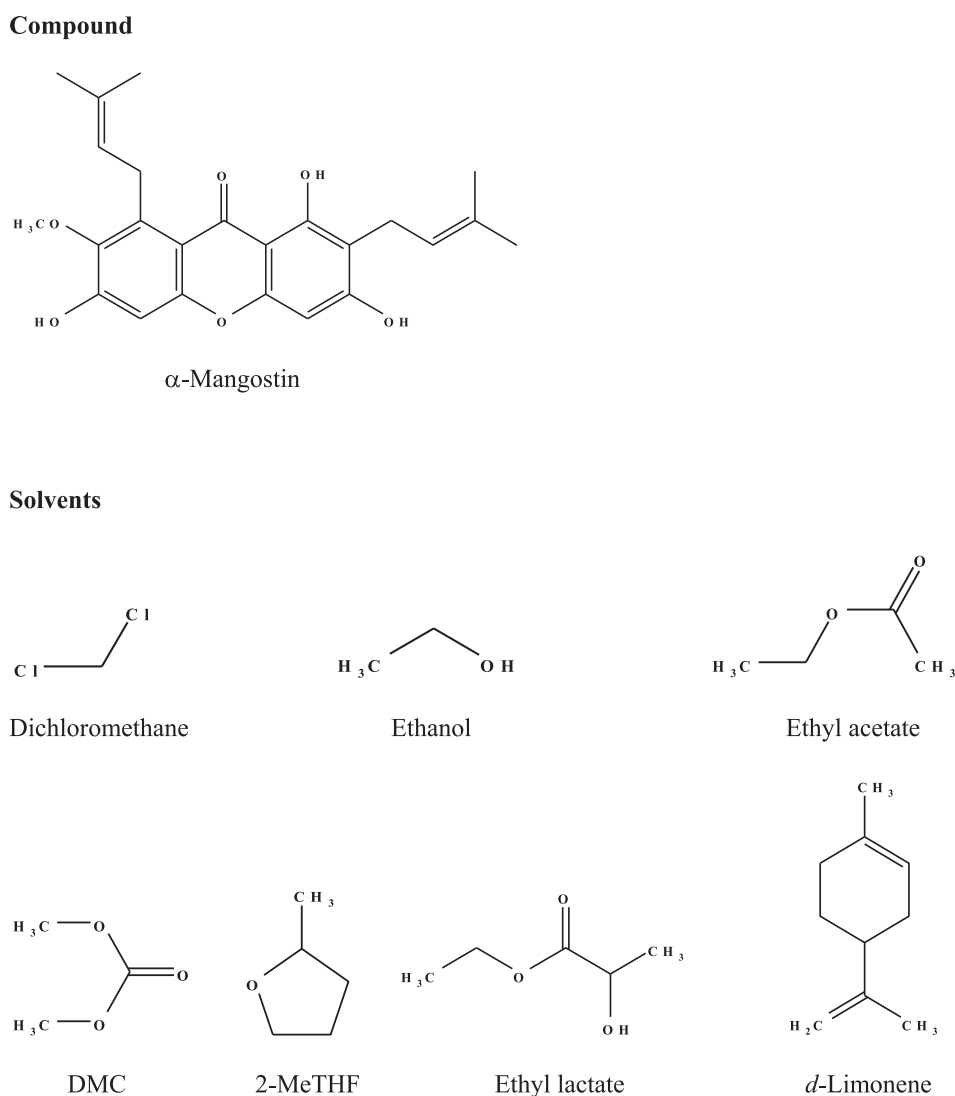
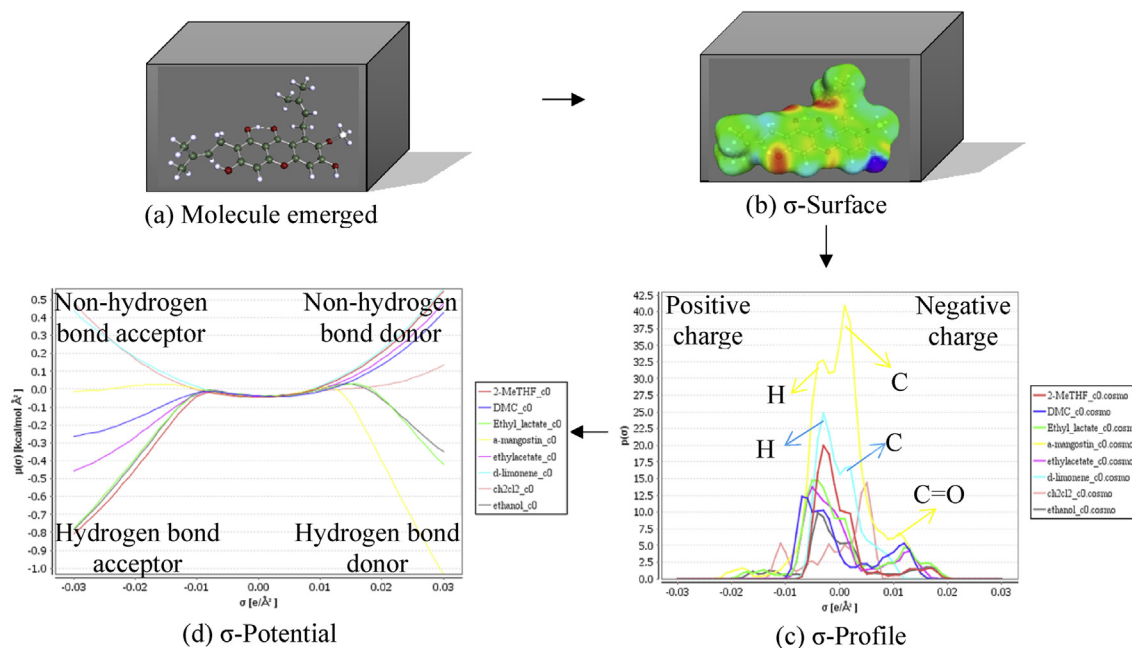


Fig. 2. Structures of  $\alpha$ -mangostin and solvents.

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**Fig. 3.** Step calculation with COSMO-RS: (a) molecule emerged; (b)  $\sigma$ -surface; (c) energies of local surface interactions between  $\sigma$ -profiles of  $\alpha$ -mangostin and solvents; (d)  $\sigma$ -potentials of  $\alpha$ -mangostin and solvents.

molecular interactions that were based on the obtained  $\sigma$ -profile was then used to calculate the chemical potential of the surface segment ( $\sigma$ -potential) as shown in Fig. 3d using COSMOthermX program (version C30 release 13.01). The  $\sigma$ -potential described the likeliness of the compound being able to interact with the solvents with polarity and hydrogen bonds (Fig. 3d). The part of the negative charge of the molecule was located on the right side (acceptor hydrogen bonds) with positive  $\sigma$  values while the part of the positive charged was located on the left side (donor hydrogen bonds) with negative  $\sigma$  values. Generally, the region  $\sigma \pm 0.01 \text{ e}/\text{Å}^2$  was considered to be non-polar or weakly polar. The  $\sigma$ -profile and the  $\sigma$ -potential were used for interpreting the affinity of the solvent for surface polarity, to understand the interaction between the compound and solvent in mixed state and to estimate the thermodynamic properties of the system.

In addition, the COSMOthermX also calculated the relative solubility between the solid compound and the liquid solvent in terms of the logarithm of the solubility in mole fractions ( $\log_{10}(x_{\text{solub}})$ ). The logarithm of the best solubility was set to 0 and all other solvents were given relative to the best solvent. The calculation of  $\alpha$ -mangostin and solvents was performed at the boiling point of each solvent.

### 2.5. Solid-liquid extraction

The dried powders of *G. mangostana* pericarp (2 g) were sepa-

The yield of  $\alpha$  – mangostin = weight of  $\alpha$  – mangostin/weight of *G. mangos tan a powder*

(4)

rately extracted with DMC, 2-MeTHF, ethyl lactate, ethyl acetate, ethanol, *D*-limonene and dichloromethane (20 mL) under a conventionally heated reflux at the boiling point of each solvent (Table 1) for 1 h. The extracts were then filtered and evaporated

using a rotary vacuum evaporator. The extracts were subjected to HPLC analysis for their  $\alpha$ -mangostin content. The extraction yields were also recorded. The experiments were performed in triplicate.

### 2.6. Quantitative analysis of $\alpha$ -mangostin by HPLC

HPLC analysis of  $\alpha$ -mangostin was based on the method described by Yodhnu, Sirikatitham, and Wattanapiromsakul (2009) with minor modification. HPLC analysis was carried out using a binary HPLC pump and autosampler (Waters e2695, USA) equipped with a photodiode array detector (Waters 2998, USA). Separation was achieved at 25 °C on a Lichrocart® column (5  $\mu\text{m}$ , 250 mm length  $\times$  4 mm internal diameter). The mobile phase consisted of acetonitrile and 0.2% formic acid in an aqueous solution (70:30; v/v), and was pumped at a flow rate of 1 mL/min. The sample injection volume was 20  $\mu\text{L}$  and the quantitative detection was by UV at a wavelength of 240 nm. The calibration curve of  $\alpha$ -mangostin were established using standard  $\alpha$ -mangostin (concentration range was 12.5, 25, 50, 100 and 200  $\mu\text{g}/\text{mL}$ ). The amount of  $\alpha$ -mangostin exhibited linearity over the evaluated ranges, with the linear equations of  $Y = 78806X + 78806X$  ( $R^2 = 1$ ) respectively. Each calibration point was carried out in triplicate.

The yield of  $\alpha$ -mangostin was expressed as the weight of  $\alpha$ -mangostin extracted from the initial *G. mangostana* powder based on HPLC analysis, as described in Equation (4);

The selectivity of each solvent for extraction of  $\alpha$ -mangostin was expressed from the weight of the  $\alpha$ -mangostin content in the dried weight of *G. mangostana* extract, as described in Equation (5):

Selectivity = yield of  $\alpha$  – mangostin/weight of dried *G. mangos tan a* extract (5)

### 3. Results and discussion

#### 3.1. Solute–solvent solubility by HSPs

The HSP values of solute and solvents were calculated from their structures (Table 1) and used in plotting a three dimensional (3D) HSP sphere (Fig. 4a) for a theoretical characterization of the solute–solvent interaction. Generally, the optimum HSP for good solvency was characterized by the green dot in the sphere of the 3D graph. However, the result from the HSP showed that all blue dots were outside the sphere to indicate a poor solvency (Fig. 4a). The two dimensional (2D) graph in terms of  $\delta_p$  versus  $\delta_h$  was plotted for a better understanding of the solubility (Fig. 4b). Moreover, the calculated  $R_a$  of dichloromethane was less than the other selected

solvents (Table 1), showed a center closer to HSP and gave a better solubility for  $\alpha$ -mangostin than the other solvents. The RED numbers have been used to characterize and quantify the solute–solvent interactions. Table 1 represents the RED modeling optimization of solvents for the extraction of  $\alpha$ -mangostin. Every solvent showed various theoretical solubility for  $\alpha$ -mangostin, which was explained by the differences in the solvent polarities.

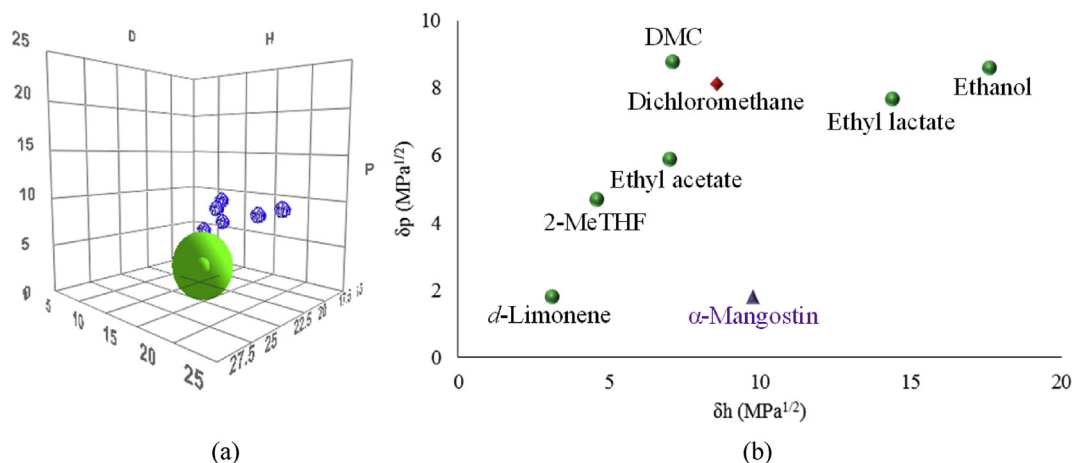
#### 3.2. COSMO-RS prediction

The COSMO-RS data was used to study and understand the potential of alternative solvents for extraction of  $\alpha$ -mangostin. Analysis of the  $\sigma$ -profile and the  $\sigma$ -potential of the components gave some important information about the molecules that could

**Table 1**  
Solvent properties.

Solvents	Dichloromethane	Ethanol	Ethyl acetate	DMC	2-MeTHF	Ethyl lactate	<i>D</i> -Limonene
<b>Physicochemical properties</b>							
Molecular weight (g/mol)	84.93	46.07	88.11	90.07	86.13	118.14	136.23
Density (g/mL)	1.25	0.78	0.90	1.00	0.86	1.04	0.83
Log P	1.19	-0.31	0.79	0.15	0.82	-0.04	4.45
Flash point (°C)	-14.1	8.90	-3.30	18.30	-11.00	48.90	48.00
<b>Toxicities</b>							
Resource	Petroleum	Cereal crop	Cereal crop	Chemical synthesis	Cereal crop	Cereal crop	Cereal crop
CMR <sup>a</sup>	2	No	No	No	No	No	No
Toxicity index	4	5	5	5	4	5	5
<b>Energy efficiency</b>							
Boiling point (°C)	40	73	77	90	80	119	175
$\delta H_{vap}$ (kJ/mol)	28.06	38.56	31.94	33.05	30.74	45.57	39.49
$C_p$ , liquid (kJ/mol*K)	0.102	0.112	0.169	0.165	0.157	0.255	0.249
Energy for evaporation of 1 kg solvent (kJ)	348.40	953.68	462.24	486.01	457.16	588.62	564.05
<b>HSPs</b>							
$\delta d$ (MPa <sup>1/2</sup> )	17.4	16.2	15.7	15.9	16.4	16.6	16.7
$\delta p$ (MPa <sup>1/2</sup> )	8.1	8.6	5.9	8.8	4.7	7.7	1.8
$\delta h$ (MPa <sup>1/2</sup> )	8.6	17.6	7.0	7.1	4.6	14.4	3.1
$\delta Total$ (MPa <sup>1/2</sup> )	21.1	25.4	18.2	19.5	17.6	23.2	17.1
$R_a$	7.78	37.86	32.71	33.75	33.45	37.00	33.59
RED	1.94	3.10	2.31	2.63	2.19	2.40	2.22
<b>COSMO-RS</b>							
Log10( $x_{solub}$ )	0.16	0	0	0	0	0	0.16

<sup>a</sup> CMR, carcinogenic, mutagenic and reprotoxic according to European legislation.



**Fig. 4.** (a) General three-dimensional Hansen solubility sphere and (b) represent two-dimensional  $\delta_p$  versus  $\delta_h$  of solvents (bio-based  $\blacklozenge$  solvents and dichloromethane) and solute ( $\blacktriangle$   $\alpha$ -mangostin).

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be used to predict possible interactions in the fluid phase. It was noticed that there was a sign inversion compared to the molecular polarity. For example, the  $\sigma$ -profile of  $\alpha$ -mangostin showed three peaks. The first peak in the negatively charged part resulted from the polarized hydrogen atoms, the second peak at a  $\sigma$  of 0.005 e/A<sup>2</sup> resulted from the carbon atoms and the last peak in the positive charged part resulted from the carbonyl group (Fig. 3c). As a consequence, the  $\sigma$ -potential was asymmetric due to the electron pairs from the carbonyl group on the negative side (Fig. 3d). The  $\sigma$ -profile of *D*-limonene showed two peaks, resulting from the hydrogen atoms on the negative side and from the carbons atoms on the positive side. As a consequence, the  $\sigma$ -potential of *D*-limonene was similar to the parabola center at  $\sigma = 0$ , a characteristic of a non-polar alkane.

The solubility study between the  $\alpha$ -mangostin and solvents are written as  $\log_{10}(x_{\text{solub}})$  (Table 1).  $\alpha$ -Mangostin was more soluble in the solvents that had a  $\log_{10}(x_{\text{solub}})$  near 0. The solubility, when compared at the boiling point of each solvent,  $\log_{10}(x_{\text{solub}})$  of ethanol, ethyl acetate, ethyl lactate, 2-MeTHF and DMC were close to 0, unlike dichloromethane and *D*-limonene (Table 1). Therefore,  $\alpha$ -mangostin was more soluble in ethanol, ethyl acetate, ethyl lactate, 2-MeTHF and DMC than in dichloromethane and *D*-limonene.

### 3.3. $\alpha$ -Mangostin content and the selectivity of the solvents studied

The experimental study was performed on *G. mangostana* pericarp powders using different solvents. The content of  $\alpha$ -mangostin in each solvent was plotted in Fig. 5. All of the alternative solvents used were capable of extracting  $\alpha$ -mangostin and gave a higher content of  $\alpha$ -mangostin than for dichloromethane. The content of  $\alpha$ -mangostin extracted from the *G. mangostana* powders was in the following order: ethyl lactate > ethanol > DMC > 2-MeTHF > ethyl acetate > *D*-limonene > dichloromethane. The highest  $\alpha$ -mangostin content (4.19 g/100 g) was observed in ethyl lactate. The content was 2-fold higher than that of the dichloromethane extract (2.08 g/100 g). The  $\alpha$ -mangostin contents in the *G. mangostana* powder were different due to the difference in polarity and solubility between  $\alpha$ -mangostin and the solvents. The

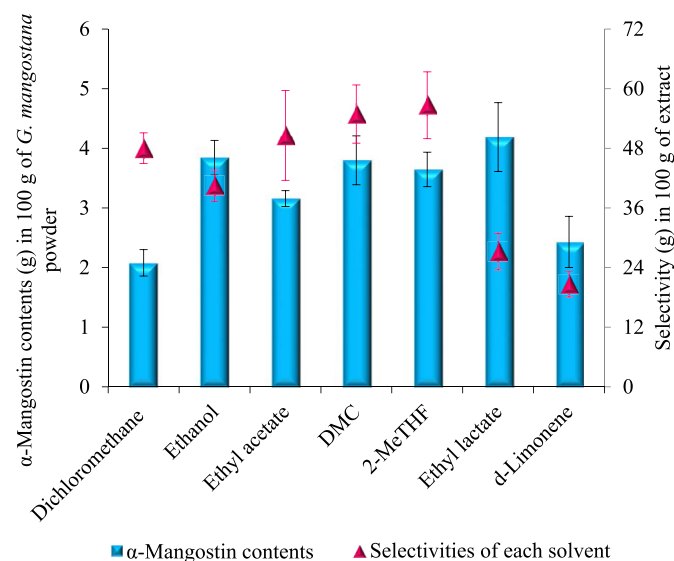


Fig. 5. Quantitative analysis of  $\alpha$ -mangostin. Data are presented as the mean values of 3 replications  $\pm$  SD.

results indicated that ethyl lactate was the most suitable solvent for  $\alpha$ -mangostin extraction.

Furthermore, the  $\alpha$ -mangostin contents in the dried extracts of each solvent were determined for their selectivity (Fig. 5). The selectivity of each solvent was found in the following order: 2-MeTHF > DMC > ethyl acetate > dichloromethane > ethanol > ethyl lactate > *D*-limonene. The highest extractability was obtained with MeTHF (56.68 g/100 g) that was much better than for dichloromethane, ethyl lactate and *D*-limonene. It was different from the previous results that had shown that ethyl lactate gave the highest content of alpha-mangostin extracted from the *G. mangostana* fruit peel powder.

Therefore, MeTHF, DMC and ethyl acetate had a better selectivity and gave a higher content of  $\alpha$ -mangostin than dichloromethane. Dichloromethane showed a better selectivity than ethanol, ethyl lactate and *D*-limonene. This was a similar finding to the previous reports on the selectivity of  $\alpha$ -mangostin extraction using ethanol and dichloromethane (Pothitirat et al., 2010).

However, on further consideration of the energy consumption by these solvents when evaporating 1 kg of solvent, these alternative solvents consumed more energy than dichloromethane (Table 1). The energy consumption was in the following order: ethanol > ethyl lactate > *D*-limonene > DMC > ethyl acetate > 2-MeTHF > dichloromethane.

MeTHF, DMC and ethyl acetate were more selective for  $\alpha$ -mangostin than ethyl lactate and dichloromethane. Dichloromethane showed a higher selectivity than ethanol, ethyl lactate and *D*-limonene. A similar, results had been previously reported using ethanol and dichloromethane (Pothitirat et al., 2010), and this supported the present finding. However, after further consideration of the energy consumption by these solvents when evaporating 1 kg of solvent, these alternative solvents required more energy than dichloromethane (Table 1).

In addition, these solvents were not carcinogenic, mutagenic, and reprotoxic (CMR) compounds. DMC was produced by green organic synthesis (Tundo & Selva, 2002), 2-MeTHF is a by-product produced from the waste of corn cobs (Antonucci et al., 2011), ethyl acetate is found in rum and grapes damaged by hail or it was produced from ethanol and acetic acid (Filly et al., 2015). Fortunately, all these solvents also showed a higher selectivity for extracting  $\alpha$ -mangostin than the other solvents. Therefore, DMC, 2-MeTHF and ethyl acetate can be promising alternative solvents for the extraction of  $\alpha$ -mangostin.

### 3.4. Comparisons of the experimental, Hansen and COSMO-RS studies

The experimental data were compared with the results from the predictive modeling: Hansen and COSMO-RS as shown in Table 2.

The Hansen method described dichloromethane as a good solvent for  $\alpha$ -mangostin extraction, but not according to the experimental data (Table 1). This was mainly because the HSP was reported to have defects in its simulation when studied with non-standard solvents with large molecules or complex multi-component systems (Filly et al., 2014).

On other hand, the COSMO-RS method indicated that ethyl lactate, ethanol, DMC, 2-MeTHF and ethyl acetate had a good solubility for  $\alpha$ -mangostin according to the  $\log_{10}(x_{\text{solub}})$  values at each boiling point. It was of interest that the results of the COSMO-RS model correlated well and were consistent with the experimental data. Therefore, the COSMO-RS model was valid for the screening of solvents for  $\alpha$ -mangostin extraction and allowed for the emerging of bio-based molecules.

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Table 2

Comparison of Hansen, COSMO-RS and experimental studies.

Solvents	Hansen (RED <sup>a</sup> )	COSMO-RS (log10(x <sub>solub</sub> ))	Experiment ( $\alpha$ -mangostin content)
Dichloromethane	Medium	Low	Low
Ethanol	Low	High	High
Ethyl acetate	Low	High	High
DMC	Low	High	High
2-MeTHF	Low	High	High
Ethyl lactate	Low	High	High
D-Limonene	Low	Low	Low

<sup>a</sup> RED, relative energy difference.

#### 4. Conclusion

Application of the predictive COSMO-RS model was better suited than the Hansen model for selecting alternative solvents for extracting  $\alpha$ -mangostin. The experimental study confirmed that the COSMO-RS simulation was more suitable for screening the solvent for complex molecules. From this study it was concluded that DMC, 2-MeTHF and ethyl acetate were promising green solvents for the extraction of bioactive compounds involved with the pharmaceutical industry, with high extraction efficacy, high selectivity, the saving of time and energy as well as being an economically viable and environmentally friendly approach.

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